

# 6.1

## ORGANIZING THE ELEMENTS

### Section Review

#### Objectives

- Explain how elements are organized in a periodic table
- Compare early and modern periodic tables
- Identify three broad classes of elements

#### Vocabulary

- periodic law
- metals
- nonmetals
- metalloids

#### Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, element, phrase, or number.

- Chemists used the \_\_\_\_\_ 1 \_\_\_\_\_ elements to sort them into groups. \_\_\_\_\_ 5 \_\_\_\_\_
- The periodic table organizes the elements into vertical \_\_\_\_\_ 2 \_\_\_\_\_ \_\_\_\_\_ 6 \_\_\_\_\_ and horizontal \_\_\_\_\_ 3 \_\_\_\_\_ in order of increasing \_\_\_\_\_ 4 \_\_\_\_\_. The table is constructed so that elements that have similar chemical properties \_\_\_\_\_ 7 \_\_\_\_\_ \_\_\_\_\_ 8 \_\_\_\_\_ have a high fusion or boiling point and \_\_\_\_\_ 9 \_\_\_\_\_ at room temperature. Elements with properties that are similar to those of metals and nonmetals are called \_\_\_\_\_ 10 \_\_\_\_\_. Across the periodic table, the properties of elements become \_\_\_\_\_ 11 \_\_\_\_\_ variable and \_\_\_\_\_ 12 \_\_\_\_\_ nonmetallic. \_\_\_\_\_ 13 \_\_\_\_\_

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