

The mixture of the hydrocarbon and oxygen is sparked so that a complete combustion reaction occurs, producing  $\text{CO}_2(g)$  and  $\text{H}_2\text{O}(g)$ . The partial pressures of these gases at 127°C are 0.600 atm for  $\text{CO}_2(g)$  and 0.800 atm for  $\text{H}_2\text{O}(g)$ . There is  $\text{O}_2(g)$  remaining in the container after the reaction is complete.

- (a) Use the partial pressures of  $\text{CO}_2(g)$  and  $\text{H}_2\text{O}(g)$  to calculate the partial pressure of the  $\text{O}_2(g)$  consumed in the combustion.
- (b) On the basis of your answers above, write the balanced chemical equation for the combustion reaction and determine the formula of the hydrocarbon.



$$0.6 \text{ atm CO}_2 \cdot \frac{1 \text{ O}_2}{1 \text{ CO}_2} = 0.6 \text{ atm O}_2$$

$$0.8 \text{ atm H}_2\text{O} \cdot \frac{1 \text{ O}_2}{2 \text{ H}_2\text{O}} = 0.4 \text{ atm O}_2$$

1.0 atm O<sub>2</sub>

# 2012 Ap Chemistry Response Answers

**IM Harris**



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