

Bubling Point Determination Pre-Lab Questions

1. Define boiling point.

The boiling point of a substance is defined as the temperature at which the pressure of the vapor of the substance becomes equal to the prevailing external pressure above the substance

2. How will an increase in pressure affect the boiling point of a substance?

The greater the pressure, the more energy required for liquids to boil, and the higher the boiling point.

3. What two reference points will you use to calibrate your temperature.

The freezing point of water (0°C) and the boiling point of water (100°C), but the temperatures depend on the atmospheric pressure

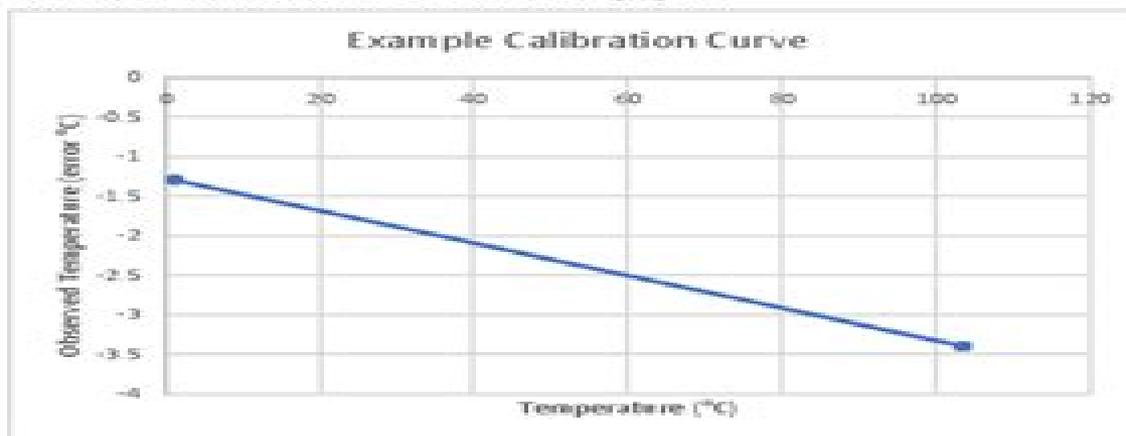
4. After finishing your experiment, what is the proper procedure for disposing the remaining unknown in your small test tube?

Discard the unknown liquid from the small test tube into the bottle labeled waste experiment 1 boiling point unknowns

5. Look up in your textbook or another reference source the boiling points of the three different unknowns used in this experiment.

methanol = 64.7°C ethanol = 78.4°C isopropyl alcohol = 82.5°C

6. Draw the calibration curve described on page 6.



Boiling Point Lab Questions

Annelies Wilder-Smith



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