

| Question number | Answer | Marks | Guidance | | | | | | | | | | | | | | | |
|----------------------|---|----------|-----------|----------|-----------|-----------------|-----------------|----|----|----------------------|----|------------------|----|----|----|----|----|--|
| 1 (a) | C | B1 | | | | | | | | | | | | | | | | |
| 1 (b) | B | B1 | | | | | | | | | | | | | | | | |
| 1 (c) | D | B1 | | | | | | | | | | | | | | | | |
| 2 (a) | 38 | B1 | | | | | | | | | | | | | | | | |
| 2 (b) | 32 | B1 | | | | | | | | | | | | | | | | |
| 2 (c) | 5 | B1 | | | | | | | | | | | | | | | | |
| 3 (a) (i) | 23 | B1 | | | | | | | | | | | | | | | | |
| 3 (a) (ii) | $^{56}_{26}\text{Fe}$ | B2 | | | | | | | | | | | | | | | | |
| 3 (b) | $^{23}_{11}\text{Na}$ | B2 | | | | | | | | | | | | | | | | |
| 3 (c) | $^{88}_{38}\text{Sr}$ | B2 | | | | | | | | | | | | | | | | |
| 4 (a) | Isotopes are atoms of the same element with different numbers of neutrons and different masses. | B1 | | | | | | | | | | | | | | | | |
| 4 (b) | Relative isotopic mass is the mass of an isotope relative to one-twelfth of the mass of an atom of carbon-12. | B2 | | | | | | | | | | | | | | | | |
| 4 (c) | 71 | B1 | | | | | | | | | | | | | | | | |
| 4 (d) | <table border="1"> <thead> <tr> <th></th> <th>protons</th> <th>neutrons</th> <th>electrons</th> <th>charge</th> </tr> </thead> <tbody> <tr> <td>^{17}O</td> <td>8</td> <td>9</td> <td>10</td> <td>2-</td> </tr> <tr> <td>^{27}Al</td> <td>13</td> <td>14</td> <td>10</td> <td>3+</td> </tr> </tbody> </table> | | protons | neutrons | electrons | charge | ^{17}O | 8 | 9 | 10 | 2- | ^{27}Al | 13 | 14 | 10 | 3+ | B2 | |
| | protons | neutrons | electrons | charge | | | | | | | | | | | | | | |
| ^{17}O | 8 | 9 | 10 | 2- | | | | | | | | | | | | | | |
| ^{27}Al | 13 | 14 | 10 | 3+ | | | | | | | | | | | | | | |
| 5 (a) | Relative atomic mass is the weighted mean mass of an atom of an element relative to one-twelfth of the mass of an atom of carbon-12. | B3 | | | | | | | | | | | | | | | | |
| 5 (b) | $A_r = 32.09$ | B2 | | | | | | | | | | | | | | | | |
| 5 (c) | <table border="1"> <thead> <tr> <th></th> <th>protons</th> <th>neutrons</th> <th>electrons</th> </tr> </thead> <tbody> <tr> <td>^{32}S</td> <td>16</td> <td>16</td> <td>16</td> </tr> <tr> <td>$^{32}\text{S}^{2-}$</td> <td>16</td> <td>16</td> <td>18</td> </tr> </tbody> </table> | | protons | neutrons | electrons | ^{32}S | 16 | 16 | 16 | $^{32}\text{S}^{2-}$ | 16 | 16 | 18 | B1 | | | | |
| | protons | neutrons | electrons | | | | | | | | | | | | | | | |
| ^{32}S | 16 | 16 | 16 | | | | | | | | | | | | | | | |
| $^{32}\text{S}^{2-}$ | 16 | 16 | 18 | | | | | | | | | | | | | | | |
| 6 (a) | $3\text{Mg}(\text{s}) + \text{N}_2(\text{g}) \rightarrow \text{Mg}_3\text{N}_2(\text{s})$ 1 mark for correct species and state symbols 1 mark for balance | B1 × 2 | | | | | | | | | | | | | | | | |
| 6 (b) | $\text{Ca}(\text{s}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow \text{Ca}(\text{OH})_2(\text{aq}) + \text{H}_2(\text{g})$ 1 mark for correct species and state symbols 1 mark for balance | B1 × 2 | | | | | | | | | | | | | | | | |
| 6 (c) | $6\text{NaOH}(\text{aq}) + \text{Fe}_2(\text{SO}_4)_3(\text{aq}) \rightarrow 2\text{Fe}(\text{OH})_3(\text{s}) + 3\text{Na}_2\text{SO}_4(\text{aq})$ 1 mark for correct species and state symbols 1 mark for balance | B1 × 2 | | | | | | | | | | | | | | | | |

Answers To Unit 2 As Level Chemistry

Lei Shi



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