NAME: KEY DATE: PERIOD:

AP CHEMISTRY MULTIPLE CHOICE QUESTIONS
Unit 8 – Introduction to Equilibrium

1. $N_2O_4(g) \rightleftharpoons 2\,NO_2(g) + \frac{1}{2}O_2(g)$

The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?

$O_2(g) + 4\,NO_2(g) \rightleftharpoons 2\,N_2O_4(g)$ Reversed & Doubled

(A) $(95)^2$

(B) 95

(C) $(95)^{\frac{1}{2}}$

(D) $\frac{1}{95}$

(E) $\frac{1}{(95)^2}$

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products? increase "P"

(A) $SrCO_3(s) \rightleftharpoons SrO(s) + CO_2(g)$

(B) $C(s) + CO_2(g) \rightleftharpoons 2\,CO(g)$

(C) $CO(g) + H_2O(g) \rightleftharpoons CO_2(g) + H_2(g)$

(D) $PCl_3(g) \rightleftharpoons PCl_3(g) + Cl_2(g)$

(E) $2\,NO_2(g) \rightleftharpoons N_2O_4(g)$

shift towards fewer # moles gas

3. $C_2H_4(g) + H_2(g) \xrightarrow{\text{platinum}} C_2H_6(g)$

$C_2H_4(g)$ is reduced by $H_2(g)$ in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

I. Changes in the partial pressure of $H_2(g)$
II. Changes in the particle size of the platinum catalyst
III. Changes in the temperature of the reaction system

(A) III only
(B) I and II only
(C) I and III only
(D) II and III only
(E) I, II, and III

4. $X(g) + Y(g) \rightleftharpoons 2\,Z(g)$

When 4.00 mol each of $X(g)$ and $Y(g)$ are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of $Z(g)$ is produced. What is the value of the equilibrium constant, K_c ?

(A) 3
(B) 6
(C) 8
(D) 16
(E) 36

	X	+ Y	⇌	2Z
	4M	4M		0
	−x	−x		+2x
	1M	1M		6M

$2x = 6M$
$x = 3M$

So: $K_c = \frac{[Z]^2}{[X][Y]} = \frac{6^2}{1 \times 1} = 36$

5. $2\,BaO_2(s) \rightleftharpoons 2\,BaO(s) + O_2(g)$ ← Needs to shift left

$\Delta H^\circ = 162\ kJ/mol_{rxn}$ endothermic

A sealed rigid vessel contains $BaO_2(s)$ in equilibrium with $BaO(s)$ and $O_2(g)$ as represented by the equation above. Which of the following changes will increase the amount of $BaO_2(s)$ in the vessel?

(A) Removing a small amount of $O_2(g)$ Rx
(B) Removing a small amount of $BaO(s)$
(C) Adding He gas to the vessel X
(D) Lowering the temperature

6. $CO(g) + 2\,H_2(g) \rightleftharpoons CH_3OH(g)$ + heat

The synthesis of $CH_3OH(g)$ from $CO(g)$ and $H_2(g)$ is represented by the equation above. The value of K_c for the reaction at 483 K is 14.5.

increase heat shifts left, decrease Kc

Which of the following explains the effect on the equilibrium constant, K_c, when the temperature of the reaction system is increased to 650 K?

(A) K_c will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.
(B) K_c will increase because there are more reactant molecules than product molecules.
(C) K_c will decrease because the reaction is exothermic.
(D) K_c is constant and will not change.

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Robert R. Redfield

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