

## 9.6 MC Questions



In a blast furnace, the reaction represented above occurs, producing  $\text{Fe}_{(l)}$  from its ore,  $\text{Fe}_2\text{O}_{3(s)}$ .

The reaction is thermodynamically favorable and based on coupling the two reactions represented below.



Which of the following identifies a limitation in how the representations above describe a system of coupled reactions?

- A None of the equations are balanced.
- B The values of  $\Delta S^\circ$  for each reaction are not shown.
- C The values of  $\Delta G^\circ$  for each reaction are not shown.
- D The coupled reactions lack a common intermediate species.

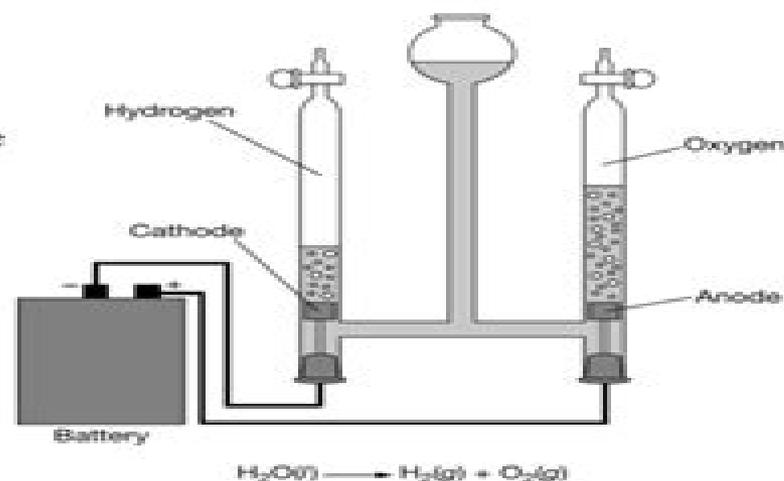
2. The reaction in which  $\text{H}_2\text{O}_{(l)}$  is decomposed into  $\text{H}_{2(g)}$  and  $\text{O}_{2(g)}$  is thermodynamically unfavorable ( $\Delta G^\circ > 0$ ).



However, an electrolytic cell, such as the one represented opposite, can be used to make the reaction occur.

Which of the following identifies a flaw in the representation?

- A Oxidation is occurring at the anode instead of at the cathode.
- B The equation for the reaction is not correctly balanced.
- C Electrical energy is needed for the reaction to proceed.
- D The relative volumes of collected gas shown in each tube are incorrect.



# Ap Electrochemistry Problems

**L Reisser**



## **Ap Electrochemistry Problems:**

This book delves into Ap Electrochemistry Problems. Ap Electrochemistry Problems is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Ap Electrochemistry Problems, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Ap Electrochemistry Problems
- Chapter 2: Essential Elements of Ap Electrochemistry Problems
- Chapter 3: Ap Electrochemistry Problems in Everyday Life
- Chapter 4: Ap Electrochemistry Problems in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Ap Electrochemistry Problems. This chapter will explore what Ap Electrochemistry Problems is, why Ap Electrochemistry Problems is vital, and how to effectively learn about Ap Electrochemistry Problems.

3. In chapter 2, this book will delve into the foundational concepts of Ap Electrochemistry Problems. The second chapter will elucidate the essential principles that must be understood to grasp Ap Electrochemistry Problems in its entirety.

4. In chapter 3, this book will examine the practical applications of Ap Electrochemistry Problems in daily life. The third chapter will showcase real-world examples of how Ap Electrochemistry Problems can be effectively utilized in everyday scenarios.

5. In chapter 4, this book will scrutinize the relevance of Ap Electrochemistry Problems in specific contexts. This chapter will explore how Ap Electrochemistry Problems is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, this book will draw a conclusion about Ap Electrochemistry Problems. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Ap Electrochemistry Problems.

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