

Chapter 10: Basic Review Worksheet

1. How is the concept of energy defined?
2. As a ball rolls down a hill _____ energy decreases, _____ energy increases, and _____ energy remains the same.
3. What is meant by the term *state function*? Provide an example.
4. What does temperature measure?
5. Explain what is meant by the terms *exothermic* and *endothermic*.
6. Define the terms *system* and *surroundings*.
7. What are some common units of energy and how are these units defined?
8. For each of the following amounts of energy, perform the indicated conversion of units.
 - a. 459 J to calories
 - b. 55.33 kJ to joules
 - c. 84.1 kJ to kilocalories
9. Calculate ΔH for each of the following cases:
 - a. $q = 34 \text{ J}$, $w = -22 \text{ J}$
 - b. $q = -28 \text{ J}$, $w = -25 \text{ cal}$
 - c. $q = -15 \text{ J}$, $w = +12 \text{ J}$
10. For an endothermic process, q is reported with a _____ sign, and for an exothermic process, q is reported with a _____ sign.
11. What is meant by the term *specific heat capacity* of a material?
12. Calculate the mass (in grams) of each of the following substances that could be warmed over the indicated temperature range by application of exactly 1.0 kJ of energy.
 - a. water, from 15°C to 42°C
 - b. iron, from 25°C to 125°C
 - c. carbon, from -10°C to 47°C
13. What is meant by the term *enthalpy*?
14. Why does the quality of energy decrease when we use it?
15. What is the driving force for every exothermic process?

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**William J. Van Nostran, William Van
Nostran**



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