

## ABAP Programming Guidelines

### Separation of Concerns

#### Rule

Follow the separation of concerns principle. Model your applications strictly as service-orientated applications. It is especially important that you separate the logic of the application layer from the logic of the presentation layer, the persistency layer and the logic for communication with external systems. Encapsulate the repository objects of the individual concerns in separate packages.

#### Details

The SoC principle identifies the parts of an application with a specific purpose and encapsulates these parts in closed units. These units only communicate with each other using specified interfaces. Thanks to this principle, the software - which would have otherwise been overcomplicated - is divided up into manageable components. As a result, the software is:

- more stable
- easier to understand
- easier to reuse
- easier to transport
- easier to maintain
- easier to test

Regarding the last point, it would even be true to say that following the SoC principle is a prerequisite for executing isolated, automated [module tests](#).

#### Bad example

The following graphic shows two examples of an ABAP application that do not follow the SoC principle.

# Abap Programming Guidelines

**Arturo Cuomo**



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web for anthropologists both smartphones and obsidian blades are forms of material culture produced through specialized technologies that is technology refers to the knowledge and skills required to make blades phones and other objects of

material culture material culture is not just found in museums of course material culture is all around

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web jan 26 2016 the solubility product or ksp is a way of showing how effective a chemical is at dissolving into an aqueous solution the higher the ksp the more the chemical will dissolve into that solvent for this lab calcium hydroxide is created then ionized  $\text{Ca(OH)}_2$  when calcium is put in water reacts as follows  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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web k sp of calcium hydroxide 1 author j m mccormick last update ocober 13 2013 introduction an equilibrium constant  $K$  is related to  $\Delta G$  for a process through eqn 1 and as long as  $\Delta H$  and  $\Delta S$  are independent of temperature eqn 1 can be combined with eqn 2 to determine  $\Delta H$  and  $\Delta S$  from the temperature dependence of  $K$

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web sep 18 2020 in this experiment we will titrate a saturated calcium hydroxide solution with a standardized hcl solution we will determine the hydroxide ion concentration for the saturated calcium hydroxide

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web your primary objective in this experiment is to test a saturated solution of calcium hydroxide and use your observations and measurements to calculate the ksp of the compound you will do this by titrating the prepared  $\text{Ca(OH)}_2$  solution with a standard hydrochloric acid solution

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web in today s experiment you will determine the solubility product ksp of calcium hydroxide  $\text{Ca(OH)}_2$  by measuring the concentration of  $\text{Ca(OH)}_2$  in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of  $\text{Ca(OH)}_2$  is represented in equation form as shown below  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a ph regulating agent notice that calcium hydroxide is divalent and thus has twice

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the

solution  $\text{Ca}(\text{OH})_2$  à  $\text{Ca}^{2+}$   $5 \times 10^{-2}$  m calcium ions  $5 \times 10^{-2}$  m  $2 \text{OH}^-$  m hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the ksp of a slightly soluble basic salt and how solubility changes with temperature ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is  $\text{Ca}(\text{OH})_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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web question experiment 8 determining ksp for calcium hydroxide post lab report sheet data table 2 points trial 1 trial 2 hci molarity 0.050 m volume of filtered  $\text{Ca}(\text{OH})_2$  solution ml 5.00 5.00 3.71 13.75 equivalence point volume ml oh in saturated  $\text{Ca}(\text{OH})_2$  m ca in saturated  $\text{Ca}(\text{OH})_2$  m ksp of  $\text{Ca}(\text{OH})_2$  average ksp of  $\text{Ca}(\text{OH})_2$

chemistry 1210 lab 11 experiment 11 determining the ksp of calcium - Sep 08 2023

web the chemical equation used was  $\text{Ca}(\text{OH})_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  the solubility product was determined to be ksp  $\text{Ca}(\text{OH})_2$  from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as  $\text{Ca}_3(\text{PO}_4)_2$  is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore  $K = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$

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web lab report 8 determinin of calcium hydroxide the purpose of this was to titrat e a saturated  $\text{Ca}(\text{OH})_2$  soultion with a standard hcl solution and find the ksp or the oh concentration was f ound and then the concentration was found and u for this experiment a titration of calcium hydroxide with hydrochlo

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for ca was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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