Cracking the Periodic Table Code

Why aren't the elements listed in alphabetical order?

Why?

At charge go, the periodic table is a bit odd. It's not square. Large portions of the table appear to be missing at the top. It's not organized alphabetically so elements can be found easily. But to a chemist, the periodic table is a very powerful tool. The periodic table is organized by properties, both chemical and physical. Those properties relate to the electronic structure of the atoms of each element. In today's activity you are going to study how the ground state electron configurations and structure of atoms are related to the shape and organization of the periodic table.

Model 1 – Blank Periodic Table

	s-block																	
1	$1s^1$													p-block				
2	$2s^1$	$2s^2$																
3	$3s^1$	$3s^2$					d-block											
4	$4s^1$	$4s^2$		$3d^1$	$3d^2$	$3d^3$		$3d^5$	$3d^6$	$3d^7$								
5	$5s^1$	$5s^2$		$4d^1$	$4d^2$	$4d^3$		$4d^5$	$4d^6$	$4d^7$					$5p^1$		$5p^3$	$5p^4$
6	$6s^1$	$6s^2$	*												$6p^1$		$6p^3$	$6p^4$
7	$7s^1$	$7s^2$	#															

f-block

*				$4f^1$	$4f^2$	$4f^3$					$4f^{11}$	$4f^{12}$	$4f^{13}$	
#				$5f^1$	$5f^2$	$5f^3$					$5f^{11}$	$5f^{12}$	$5f^{13}$	

1. Obtain a card with electron configurations for your group as assigned by your teacher.
   a. Record your team number ___1___ and team name _____*Alkali metals*_____.
   b. For each element on your card, write the last orbital notation appearing at the end of the configuration (the underlined portion).

   Answers will vary for each team. The answers for Team 1 are shown below.

   H $1s^1$ *Na* $3s^1$ *Rb* $5s^1$

   Li $2s^1$ *K* $4s^1$ *Cs* $6s^1$

2. What is similar about the last orbital notation appearing at the end of the configuration for each element in your set?

   Answers will vary for each team. For Team 1, the configuration always ends in s^1.

Cracking The Periodic Table Code Worksheet

Kathleen Armour

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