

unit 1 : the chemistry of life

Bonds

ionic bonds

- involving a transfer of valence electrons
- usually between a metal and a nonmetal
- strongest type of bond
- high melting and boiling points
- highly soluble in water

covalent bonds

- involving sharing of valence electrons
- usually between two nonmetals
- stronger than hydrogen bonds
- low melting and boiling points
- not very soluble in water

hydrogen bonds

- weak bonds between polar molecules
- formed between a hydrogen atom bonded to a highly electronegative atom (O, N, F) and another electronegative atom
- responsible for many of the unique properties of water



Carbon - is the basis of organic compounds
 - has four valence electrons (4 bonds)

Functional Groups - groups of atoms that are attached to a carbon atom and determine its chemical behavior

Steps: determine function

FUNCTIONAL GROUPS



Water

- Highly cohesive (hydrogen bonds)
- Highly adhesive (hydrogen bonds)
- Highly specific heat (hydrogen bonds)
- Highly density (hydrogen bonds)
- Highly solvent (hydrogen bonds)



Hydrogen Bonding

- Collectively, hydrogen bonds provide enough strength to hold water molecules together
- Adhesion - water sticks to other polar surfaces (hydrogen bonds)
- Cohesion - water sticks to other water molecules (hydrogen bonds)
- High specific heat - it takes a long time to heat up (hydrogen bonds)
- Density - ice is less dense than liquid water (hydrogen bonds)
- Solvent - water is the universal solvent (hydrogen bonds)



CARBOHYDRATES - CxHyOx

- Monomers: glucose, fructose, galactose
- Polymers: starch, glycogen, cellulose
- Linkages: alpha-1,4, alpha-1,6, beta-1,4

macromolecules

LIPIDS - CxHyOx

- Monomers: fatty acids, glycerol
- Polymers: triglycerides, phospholipids
- Linkages: ester bonds

PROTEINS - CxHyOxNz

- Monomers: amino acids
- Polymers: polypeptides, proteins
- Linkages: peptide bonds

NUCLEIC ACIDS - CxHyOxNzPz

- Monomers: nucleotides
- Polymers: DNA, RNA
- Linkages: phosphodiester bonds

AP BIO

unit 1 review

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J Rink



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