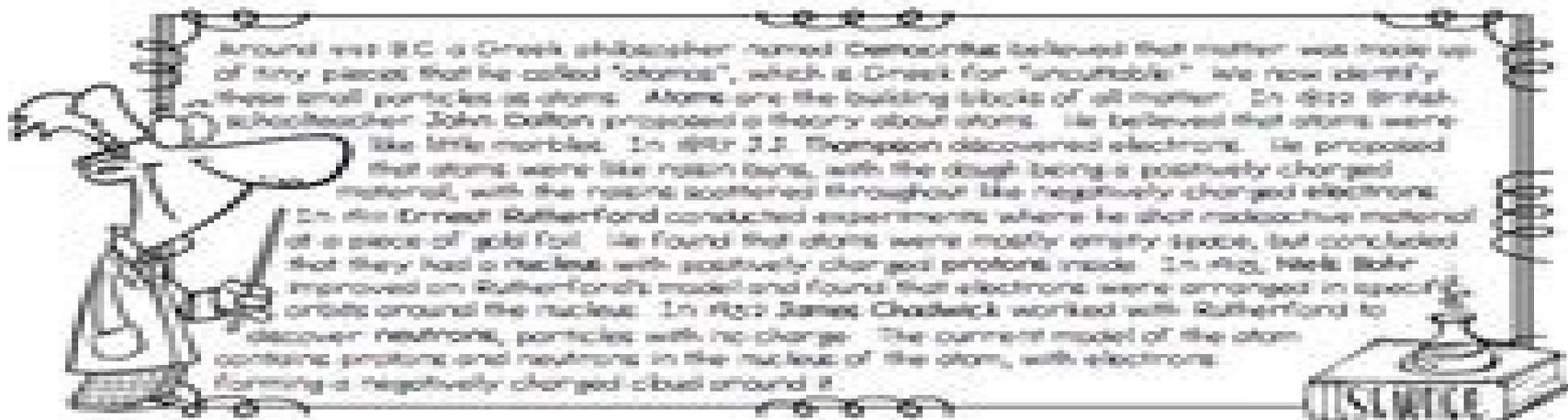


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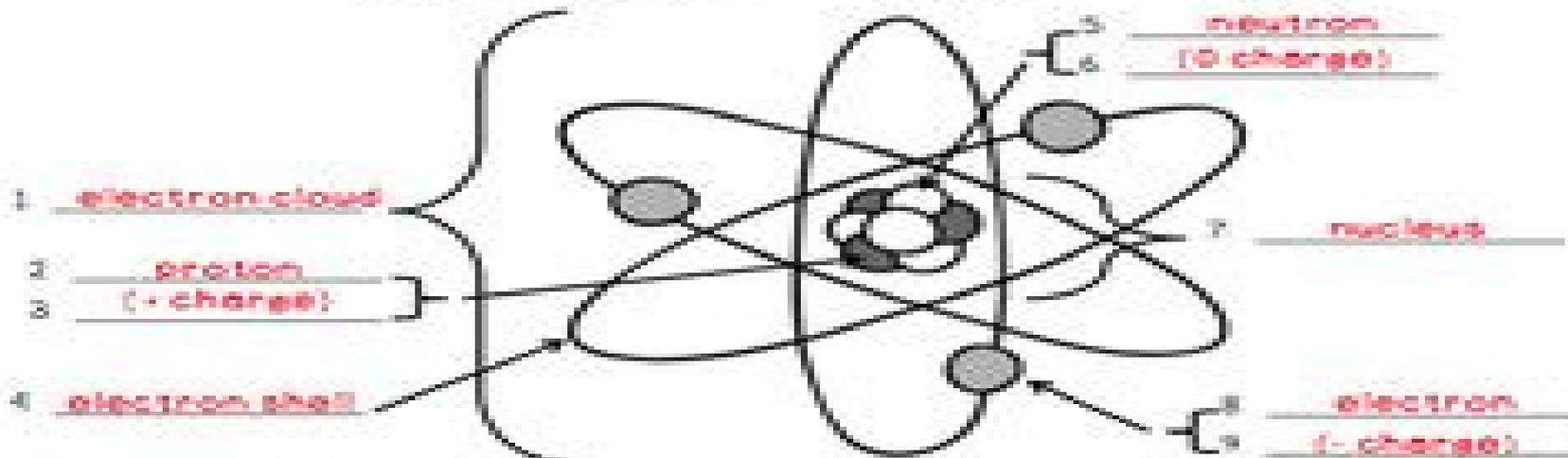


Around 450 B.C. a Greek philosopher named Democritus believed that matter was made up of tiny pieces that he called "atoms", which is Greek for "uncuttable." We now identify these small particles as atoms. Atoms are the building blocks of all matter. In 1808 British schoolteacher John Dalton proposed a theory about atoms. He believed that atoms were like little marbles. In 1897 J.J. Thompson discovered electrons. He proposed that atoms were like mush balls, with the dough being a positively charged material, with the raisins scattered throughout like negatively charged electrons. In 1911 Ernest Rutherford conducted experiments where he shot radioactive material at a piece of gold foil. He found that atoms were mostly empty space, but concluded that they had a nucleus with positively charged protons inside. In 1913, Niels Bohr improved on Rutherford's model and found that electrons were arranged in specific orbits around the nucleus. In 1932 James Chadwick worked with Rutherford to discover neutrons, particles with no charge. The current model of the atom contains protons and neutrons in the nucleus of the atom, with electrons forming a negatively charged cloud around it.

Label It

Correctly label the parts of the atom using the words and phrases below.

- Electron
(Proton)
- Electron cloud
(+ charge)
- Electron shell
(- charge)
- Neutron
(+ charge)
- Nucleus



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