

## Unit 6 Chemistry Study Guide CH 11,13 and 12.4

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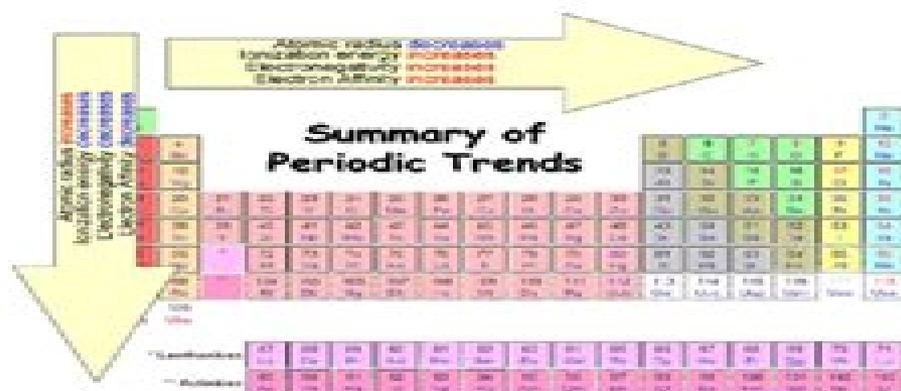
### 1. Polarizability

↑ SA/ length of , ↑ polarizability

### 2. IMF

↑ boiling/melting point, ↑ IMF

↑ vapor pressure, ↓ IMF b/c: ↑ mol solute, ↓ pressure & ↓ pressure, ↑ IMF



**Trends:**

### 1. Relate strength of intermolecular forces to state of matter

- The state of a substance depends on the balance between the *kinetic energy* of the individual particles (molecules or atoms) and the *intermolecular forces*. The kinetic energy keeps the molecules apart and moving around, and is a function of the temperature of the substance. The intermolecular forces are attractive forces that try to draw the particles together (Figure)
- As discussed previously, gasses are very sensitive to temperatures and pressure. However, these also affect liquids and solids too. Heating and cooling can change the *kinetic energy* of the particles in a substance, and so, we can change the physical state of a substance by heating or cooling it. Increasing the pressure on a substance forces the molecules closer together, which *increases* the strength of intermolecular forces.

# Chapter 11 Chemistry Study Guide

**Emilie Sanchez**



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