

CHAPTER 11 REVIEW

Ions in Aqueous Solutions and Colligative Properties

TECHNICAL

SHORT ANSWER Answer the following questions in the space provided.

1. Give the formulas of *both* ions (per cent of the correct number) of the following compounds in water:

_____ a. magnesium nitrate
 _____ b. sodium sulfate
 _____ c. calcium sulfate
 _____ d. potassium phosphate

2. List each of the following oxides as dissolved in water:

_____ a. Write the formula for magnesium oxide.
 _____ b. How many moles of ions per mole of ions will it form?
 _____ c. How many moles of ions are released into a solution with three H_2O and magnesium oxide dissolved in water?

3. Write the formula for the precipitate formed:

_____ a. when solutions of magnesium chloride and potassium phosphate are combined
 _____ b. when solutions of sodium sulfate and silver nitrate are combined

4. Write ionic equations for the dissolution of the following compounds:

a. $\text{MgSO}_4(\text{s})$

b. calcium chloride

5. a. Write the net ionic equation for the reaction that occurs when solutions of lead(II) nitrate and potassium sulfate are combined.

b. What are the spectator ions in this reaction?

Chapter 13 Review Solutions Section 1 Answers

Michael Rajnik



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