

# Equilibrium IN CLASS Practice Assessment

*Skill: The student will be able to define equilibrium, draw and explain how a catalyst changes a reaction rate graph and discuss other changes that disrupt an equilibrium system*

## 1. Complete the following:

a. What is "equal" at chemical equilibrium?

**The rates of the forward & reverse reaction**

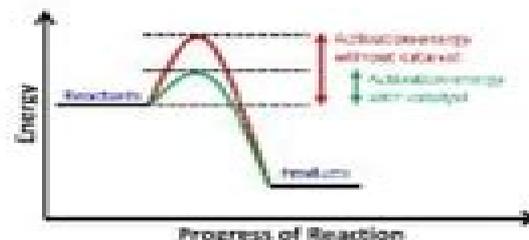
b. Provide one example of a factor that can disrupt equilibrium.

**Concentration, temp, or pressure (if gases present)**

c. Draw a new reaction pathway representing how a catalyst, like an enzyme, would affect the rate of a chemical reaction.

Provide an explanation for your new curve.

**The addition of a catalyst lowers the activation energy necessary**



*Skill: The student will be able to write the equilibrium expression, given a reaction.*

2. Write the equilibrium expression for the following reaction:



$$K = \frac{[\text{CaCl}_2][\text{CO}_2]}{[\text{HCl}]^2}$$

*Skill: The student will be able to calculate the equilibrium constant and tell if the products or reactants are favored.*

3. Calculate the equilibrium constant, for the reaction shown in Level 2, using the following equilibrium concentrations:  $\text{CaCO}_3 = 1.0\text{M}$ ,  $\text{HCl} = 0.20\text{M}$ ,  $\text{H}_2\text{O} = 0.500\text{M}$ ,  $\text{CaCl}_2 = 0.32\text{M}$  and  $\text{CO}_2 = 0.42\text{M}$

$$K = \frac{[\text{CaCl}_2][\text{CO}_2]}{[\text{HCl}]^2} \rightarrow \frac{[.32][.42]}{[.20]^2} = 3.36$$

Which are favored, reactants or products? **Forward**

*Skill: The student will be able to use Ksp correctly in an equilibrium calculation.*

4. Calculate the Ksp, for iron (III) hydroxide, if the concentration is  $4.49 \times 10^{-16}\text{M}$  at  $25^\circ\text{C}$ .



*Skill: The student will be able to apply LeChatelier's principle correctly to a given situation and explain how the equilibrium will shift to compensate for this change, referring to collision theory.*

5. To a sample of the cobalt equilibrium system studied in class, consider the following changes and say **what color** the system will turn? **Explain** using LeChatelier's principle & collision theory.



a. If magnesium chloride is dissolved in the system... **Right (Blue)**

# Concept Review Section Phase Equilibrium Answer Key

**LP Steffe**



**Concept Review Section Phase Equilibrium Answer Key:**

This book delves into Concept Review Section Phase Equilibrium Answer Key. Concept Review Section Phase Equilibrium Answer Key is a vital topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Concept Review Section Phase Equilibrium Answer Key, encompassing both the fundamentals and more intricate discussions.

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## **Table of Contents Concept Review Section Phase Equilibrium Answer Key**

1. Understanding the eBook Concept Review Section Phase Equilibrium Answer Key
  - The Rise of Digital Reading Concept Review Section Phase Equilibrium Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Concept Review Section Phase Equilibrium Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Concept Review Section Phase Equilibrium Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Concept Review Section Phase Equilibrium Answer Key
  - Personalized Recommendations
  - Concept Review Section Phase Equilibrium Answer Key User Reviews and Ratings
  - Concept Review Section Phase Equilibrium Answer Key and Bestseller Lists
5. Accessing Concept Review Section Phase Equilibrium Answer Key Free and Paid eBooks
  - Concept Review Section Phase Equilibrium Answer Key Public Domain eBooks
  - Concept Review Section Phase Equilibrium Answer Key eBook Subscription Services
  - Concept Review Section Phase Equilibrium Answer Key Budget-Friendly Options
6. Navigating Concept Review Section Phase Equilibrium Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Concept Review Section Phase Equilibrium Answer Key Compatibility with Devices
  - Concept Review Section Phase Equilibrium Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Concept Review Section Phase Equilibrium Answer Key
  - Highlighting and Note-Taking Concept Review Section Phase Equilibrium Answer Key
  - Interactive Elements Concept Review Section Phase Equilibrium Answer Key

8. Staying Engaged with Concept Review Section Phase Equilibrium Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Concept Review Section Phase Equilibrium Answer Key
9. Balancing eBooks and Physical Books Concept Review Section Phase Equilibrium Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Concept Review Section Phase Equilibrium Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Concept Review Section Phase Equilibrium Answer Key
  - Setting Reading Goals Concept Review Section Phase Equilibrium Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Concept Review Section Phase Equilibrium Answer Key
  - Fact-Checking eBook Content of Concept Review Section Phase Equilibrium Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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