

CHEMISTRY



Periodic Table of Elements

1	2											3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18								
H 1.01	He 4.00											B 10.81	C 12.01	N 14.01	O 16.00	F 18.99	Ne 20.18																		
Li 6.94	Be 9.01											Al 26.98	Si 28.09	P 30.97	S 32.06	Cl 35.45	Ar 39.95																		
Na 22.99	Mg 24.31											K 39.10	Ca 40.08											Zn 65.38	Ga 69.72	Ge 72.64	As 74.92	Se 78.96	Br 79.90	Kr 83.80					
K 39.10	Ca 40.08	Sc 44.96	Ti 47.88	V 50.94	Cr 51.99	Mn 54.94	Fe 55.85	Co 58.93	Ni 58.69	Cu 63.55	Zn 65.38	Ga 69.72	Ge 72.64	As 74.92	Se 78.96	Br 79.90	Kr 83.80	Rb 85.47	Sr 87.62	Y 88.91	Zr 91.22	Nb 92.91	Mo 95.94	Tc 98.91	Ru 101.07	Rh 101.07	Pd 106.36	Ag 107.87	Cd 112.41	In 114.82	Sn 118.71	Sb 121.76	Te 127.60	I 126.91	Xe 131.29
Rb 85.47	Sr 87.62	Y 88.91	Zr 91.22	Nb 92.91	Mo 95.94	Tc 98.91	Ru 101.07	Rh 101.07	Pd 106.36	Ag 107.87	Cd 112.41	In 114.82	Sn 118.71	Sb 121.76	Te 127.60	I 126.91	Xe 131.29	Ba 137.33	La 138.91	Ce 140.12	Pr 140.91	Nd 144.24	Pm 144.91	Sm 150.36	Eu 151.96	Gd 157.25	Tb 158.93	Dy 162.50	Ho 164.93	Er 167.26	Tm 168.93	Yb 173.05	Lu 174.97		
Cs 132.91	Ba 137.33	La* 138.91	Hf 178.49	Ta 180.95	W 183.84	Re 186.21	Os 190.23	Ir 192.22	Pt 195.08	Au 196.97	Hg 200.59	Tl 204.38	Pb 207.2	Bi 208.98	Po 209	At 210	Rn 222	Fr 223	Ra 226	Ac* 227	Th 232.04	Pa 231.04	U 238.03	Np 237	Pu 244	Am 243	Cm 247	Bk 247	Cf 251	Es 252	Fm 257	Md 258	No 259	Lr 262	
Ce 140.12	Pr 140.91	Nd 144.24	Pm 144.91	Sm 150.36	Eu 151.96	Gd 157.25	Tb 158.93	Dy 162.50	Ho 164.93	Er 167.26	Tm 168.93	Yb 173.05	Lu 174.97																						
Th 232.04	Pa 231.04	U 238.03	Np 237	Pu 244	Am 243	Cm 247	Bk 247	Cf 251	Es 252	Fm 257	Md 258	No 259	Lr 262																						



- Elemental forms:**
 - Monatomic:** Noble gases; helium, neon, argon, krypton, xenon, and radon; nonmetals.
 - Diatoms:** Gases: fluorine, chlorine, hydrogen, oxygen, and nitrogen; Liquid: bromine; Solid: iodine.
 - Molecular solids:** Sulfur.
 - Covalent solids:** Carbon (diamond, graphite).
 - Metals:** Most common type of element; high coordination numbers; very dense.
- Types of compounds:** Large variety due to ionic, covalent, and metallic bonding options.
 - Ionic:** Electron exchange.
 - Covalent:** Shared electrons.
 - Polar covalent:** Unequal sharing.
 - Ion complex:** Central ion + Solvent.
 - Transition metal compounds:**
 - Coordination compound:** Non-bonding ligands.
 - Organometallic:** Metal-C bond.
 - Complex minerals:** Metal oxides and silicates.
- Periodicity:** The chemical properties of an element depend on valence electrons.
 - Family of elements:** Column in periodic table.
 - The addition of d-orbitals produces an expanded octet.

- Ionization energies:** Noble gases are the most stable; alkali metals are the easiest to ionize; electron screening lessens Z-effect on orbital energy.
- Atomic and ionic radii:** Not real properties; derive from structural data; divide bond into radii using a consistent scheme.
 - Ionic radii:** Depends on charge, counter ion, and coordination (crystal type).
 - Atomic radii:** Based on diatomic gas data or metallic solid structural data.

Development of the Periodic Table

- Missing elements were predicted using periodicity.
- Column 1:** Alkali metals.
- Column 2:** Alkaline earth metals.
- First transition series:** Titanium-Copper.
- Column 17:** Halogens: fluorine, chlorine, bromine, iodine, and astatine.
- Column 18:** Rare gases: helium, neon, argon, krypton, xenon, and radon.
- Actinides:** Z = 90 or above; radioactive.
- Lanthanides:** f-subshell elements; starts with Z = 58.



Practical Application
The colors of fireworks depend on the elements in the explosive.



Atomic Structure

- Atomic number (Z):** The number of protons in the nucleus; for a neutral atom, Z = number of electrons; for an anion of charge -q, Z = q electrons; for a cation of charge +q, Z = q electrons.
- Mass number (A):** A = Z + N, where N = the number of neutrons in the nucleus.
- Isotopes:** Atoms with the same Z, different A; an element can have one or more isotopes; any sample contains a number of isotopes; in practical work, use the average mass number or atomic weight (weighted average of natural isotopes for a given element).



Atomic Quantum Numbers & Orbitals

- n:** principle quantum number (orbital size)
- l:** angular momentum (orbital shape)
- m_l:** magnetic (orbital direction)
- m_s:** electron spin
- For each n, the possible l values are 0, 1, ..., n - 1
- For each l, the possible m_l values are -l, 0, ..., +l
- m_s has two possible values: +1/2 and -1/2 (spin up, spin down)
- Each value of n denotes a "shell" in the atomic structure; each l denotes a subshell.
 - s-orbital: l = 0, 1 type
 - d-orbital: l = 2, 5 types
 - p-orbital: l = 1, 3 types
 - f-orbital: l = 3, 7 types

Many-Electron Atoms

Orbitals and quantum numbers derived for the hydrogen atom are used to describe all many-electron atoms and ions; a more rigorous treatment determines exact energy levels and orbital properties; the main result is that subshells have different energies; the filling of levels is guided by the Aufbau principle:



n	l	m _l	m _s	element	orbital
1	0	0	+1/2	H	1s-orbital
1	0	0	-1/2	He	spherical
2	0	0	+1/2	Li	2s-orbital
2	0	0	-1/2	Be	spherical
2	1	-1	+1/2	B	2p-orbitals
2	1	0	+1/2	C	
2	1	1	+1/2	N	hemisidal
2	1	-1	-1/2	O	p, p, p
2	1	0	-1/2	F	
2	1	+1	-1/2	Ne	

- Electron configuration:** Atomic orbital occupancy for an atom.
- Pauli exclusion principle:** Each electron has a unique set of quantum numbers; an orbital may hold up to two electrons; one with spin up, m_s = +1/2, and one with spin down, m_s = -1/2.
- Hund's rule:** Electrons fill p, d, and f subshells in a manner to maximize the number of unpaired electrons (half-filled orbitals); spin is maximized.
- Ionization potential (IP):** Energy required to remove electron from an atom or ion.
 - First IP:** Removal of first valence electron.
- Electronegativity:** Tendency of an atom to attract electrons in a chemical bond; 0-4 scale; 0 for rare gas.

College Chemistry Study Guide

Karin Nielsen-Saines



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