

Unit 8 Review

1. The balanced equation $\text{P}_4(\text{s}) + 6\text{H}_2(\text{g}) \rightarrow 4\text{PH}_3(\text{g})$ tells us that 4.0 mol H_2
 A) reacts with 2.0 mol P_4
 B) produces 8.0 mol PH_3
 C) cannot react with phosphorus
 D) produces 2.7 mol PH_3
 E) reacts with 4.0 mol P_4

2. A 1.0-mol sample of Fe_2O_3 was decomposed according to the equation:
 $2\text{Fe}_2\text{O}_3(\text{s}) \rightarrow 2\text{Fe}(\text{s}) + 3\text{O}_2(\text{g})$

How many moles of O_2 are formed assuming 100% yield?

- A) 0.0 mol
- B) 1.5 mol
- C) 1.7 mol
- D) 0.7 mol
- E) 2.0 mol

3. The rusting of iron is represented by the equation $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$. If you have a 1.00-mol sample of iron, how many moles of Fe_2O_3 will there be after the iron has rusted completely?

- A) 0.527 mol
- B) 0.750 mol
- C) 1.00 mol
- D) 1.50 mol
- E) 2.07 mol

4. Consider the following reaction, where X represents an unknown element.
 $8\text{X}_2(\text{s}) + 2\text{B}_2\text{O}_3(\text{s}) \rightarrow \text{B}_4\text{X}_6(\text{s}) + 3\text{O}_2(\text{g})$

If 175 g of X reacts completely with diboron trioxide to produce 1.43 mol of B_4X_6 , what is the identity of X?

- A) Ge
- B) Mg
- C) Si
- D) Sn
- E) C

5. Refer to the following equation:
 $4\text{NH}_3(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$

How many molecules of NO are produced when 3.02 mol of ammonia is completely reacted?

- A) 12.08
- B) 3.64×10^{23}
- C) 1.82×10^{23}
- D) 1.81
- E) none of these

6. In the reaction $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$, how many moles of ammonia would be produced from 1.00 mol of hydrogen and excess nitrogen?

- A) 1.42 mol
- B) 3.37 mol
- C) 0.727 mol
- D) 2.18 mol
- E) 0.343 mol

7. Refer to the following unbalanced equation:
 $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

What mass of oxygen (O_2) is required to react completely with 13.1 g of C_2H_6 ?

- A) 2.98×10^2 g
- B) 23.1 g
- C) 4.86 g
- D) 49.2 g
- E) 0.192 g

8. How many molecules of carbon dioxide will be formed if 7.07 g of propane is burned in the following reaction?



- A) 1.37×10^{23} molecules
- B) 1.00×10^{23} molecules
- C) 4.14×10^{23} molecules
- D) 3.10×10^{23} molecules
- E) 5.17×10^{23} molecules

9. Calculate the mass of carbon dioxide produced from 15.7 g of octane, C_8H_{18} , in the following reaction.



- A) 79.2 g CO_2
- B) 158 g CO_2
- C) 89.1 g CO_2
- D) 1.258 g CO_2
- E) 1.400 g CO_2

10. Calculate the molecules of oxygen required to react with 18.2 g of sulfur in the following reaction.



- A) 1.03×10^{23} molecules O_2
- B) 2.06×10^{23} molecules O_2
- C) 1.71×10^{23} molecules O_2
- D) 3.42×10^{23} molecules O_2
- E) 5.13×10^{23} molecules O_2

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