

Chemistry:

CHAPTER 1

- Charge of electron $\approx -1.6 \times 10^{-19}$; Avogadro's # $\approx 6 \times 10^{23}$
- Mass number (A) on top of element, atomic number (Z) below
- $E = \frac{hc}{\lambda} = -R \left[\frac{1}{n_1^2} - \frac{1}{n_2^2} \right]$; $E = qV$ (energy = charge on e- x electron volts); $c = \lambda\nu$ ($c = 3 \times 10^8$)
- **Heisenberg uncertainty** – impossible to determine momentum and position simultaneously
- **Pauli exclusion** – no electrons have same quantum #'s
- Max e- in a shell $\approx 2n^2$, max e- in subshell $\approx 4l+2$
 - Possible values for l ≈ 0 to $n-1$, possible values for ml $\approx -l$ to $+l$
- s can hold 2 e-, p ≈ 6 , d ≈ 10 , f ≈ 14
- **Aufbau principle** (building up) – lowest $n+l$ will fill first
- For anions add extra e-, for cations remove e- from subshells with highest n first
- **Hund's** – half filled orbitals with parallel spins (bus seats)
- **Paramagnetic** – unpaired parallel spin e-, attracted to magnetic field
 - **Diamagnetic** – paired e-, repelled by magnetic field
 - **Ferromagnetic** – unpaired e-, randomly oriented dipoles, become magnetized in field
- A faraday is equal to one mole of electric charge

CHAPTER 2

- **Z eff** increases from left to right, constant up and down
- **Ionic radius** is smallest near metalloid line and increases out
- F is smallest, most EN, highest IE and EA; Cs is the largest and least EN
- Complex ions from transition metals associate with water to form **hydration complexes**
- **Atomic radius** has the opposite trend of all others



CHAPTER 3

- C, N, O, F, Na, Mg always abide by octet rule; H, He, Li, Be, B can have incomplete; period 3 or higher can have expanded
- **Coordinate covalent bond** – both e- are contributed by one atom (lone pair attack, acid-base)
- Ionic \approx EN difference greater than 1.7, polar covalent \approx 0.5 to 1.7, nonpolar \approx less than 0.5
 - Ionic are stronger than covalent (which are both stronger than IM forces)
 - Van der Waals occurs more with close proximity and large molecules
- **Dipole moment** $\approx p = qd$
- Triple bonds are shorter and have more energy
- Major resonance contributor is more stable (less formal charge, negative on more EN atom)
 - **Formal charge = valence – nonbonding – ½ bonding**
- Electron geometry is when bonding and lone are equal, molecular geo is when lone repel more

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web terms in this set 11 a line on a map or chart along which there is a constant value isotherm a line on a map connecting points having the same temperature at a given time on average of a given period isobar a line on a map connecting points having the same atmospheric pressure at a given time on average of a given period

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web isotherm and isobar maps purpose the purpose of this laboratory exercise is to become familiar with how to construct

both isotherm and isobar maps vocabulary 10 points isoline isotherm isobar objectives upon completion of this laboratory exercise you will be able to 1 define the terms isoline isotherm and isobar 2 construct an

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web be interpreted as temperature therefore isotherm means same temperature isotherms drawn on temperature maps connect equal temperatures here are some helpful hints 1 isobars are drawn in 4 millibar intervals 2 always label the isobar value 3 pressures lower than the isobar value are always on one side of the isobar and higher