

balance nuclear reaction equations

aluminum-26 undergoes positron emission
 potassium-40 undergoes beta decay (electron)
 uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.
 Write the electron configuration for Cr.
 Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- | | |
|-----------------------------|---|
| a. Sulfur hexafluoride | SF_6 |
| b. Lithium nitride | Li_3N |
| c. Chromium (III) carbonate | $\text{Cr}_2(\text{CO}_3)_3$ |
| d. Tin (II) chloride | SnCl_2 |
| e. Ammonium acetate | $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$) |
| f. Mercury (I) chloride | Hg_2Cl_2 |
| g. Potassium bromate | KBrO_3 |
| h. Hydrosulfuric acid | $\text{H}_2\text{S(aq)}$ |
| i. Chloric acid | HClO_3 |
| j. Sulfurous acid | H_2SO_3 |

Name each of the following compounds:

- | | |
|------------------------------|-------------------------|
| a. CuSO_4 | Copper (II) sulfate |
| b. AlF_3 | aluminum fluoride |
| c. HI | hydrogen iodide |
| d. NO | nitrogen monoxide |
| e. $\text{H}_2\text{Se(aq)}$ | hydroselenic acid |
| f. HNO_3 | nitric acid |
| g. HNO_2 | nitrous acid |
| h. NaHSO_4 | Sodium hydrogen sulfate |

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



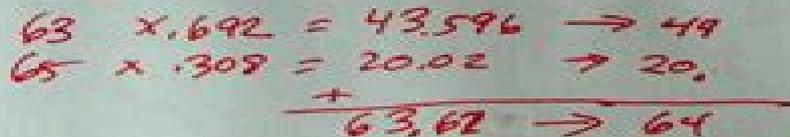
Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



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Arturo Cuomo



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 - The Rise of Digital Reading Answer Key For Study Guide Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key For Study Guide Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key For Study Guide Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key For Study Guide Chemistry
 - Personalized Recommendations

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- Answer Key For Study Guide Chemistry and Bestseller Lists
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 - Participating in Virtual Book Clubs
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 - Creating a Diverse Reading Collection Answer Key For Study Guide Chemistry
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key For Study Guide Chemistry
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