

## Balancing Chemical Reactions

Balance all of the chemical reactions below. All compounds are written correctly.

After balancing the reaction, determine the type(s) of chemical reaction(s).

TYPE(S) of RXN

- $2 \text{NH}_3 \rightarrow \text{N}_2 + 3 \text{H}_2$  Decomp
- $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$  Syn/Comb
- $\text{N}_2\text{H}_4 + \text{O}_2 \rightarrow \text{N}_2 + 2 \text{H}_2\text{O}$  S.R. (?)
- $2 \text{Al} + 3 \text{H}_2\text{SO}_4 \rightarrow 3 \text{H}_2 + \text{Al}_2(\text{SO}_4)_3$  S.R.
- $\text{P}_2\text{O}_5 + 3 \text{H}_2\text{O} \rightarrow 2 \text{H}_3\text{PO}_4$  Syn
- $\frac{1}{2} \text{C}_2\text{H}_2 + \frac{5}{2} \text{O}_2 \rightarrow \text{CO}_2 + \frac{3}{2} \text{H}_2\text{O}$  Comb.
- $\text{Fe}_2\text{O}_3 + 3 \text{H}_2 \rightarrow 2 \text{Fe} + 3 \text{H}_2\text{O}$  S.R.
- ~~4~~  $\text{FeS} + \frac{7}{2} \text{O}_2 \rightarrow \frac{2}{1} \text{Fe}_2\text{O}_3 + \frac{4}{2} \text{SO}_2$  Comb.
- ~~10~~  $\text{Cu}_2\text{S} + \frac{3}{2} \text{O}_2 \rightarrow \frac{2}{1} \text{Cu}_2\text{O} + \frac{1}{2} \text{SO}_2$  Comb.
- ~~4~~  $\text{NH}_3 + \frac{5}{2} \text{O}_2 \rightarrow \frac{4}{2} \text{NO} + \frac{6}{2} \text{H}_2\text{O}$  Comb.
- $\text{C}_3\text{H}_8 + 5 \text{O}_2 \rightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O}$  Comb
- $2 \text{AlPO}_4 + 3 \text{Na}_2\text{S} \rightarrow \text{Al}_2\text{S}_3 + 2 \text{Na}_3\text{PO}_4$  D.R.
- $\text{C}_{25}\text{H}_{52} + 38 \text{O}_2 \rightarrow 26 \text{H}_2\text{O} + 25 \text{CO}_2$  Comb
- $2 \text{FeCl}_3 + 3 \text{Na}_2\text{CO}_3 \rightarrow \text{Fe}_2(\text{CO}_3)_3 + 6 \text{NaCl}$  D.R.
- $\text{P}_2\text{O}_5 + 6 \text{H}_2\text{O} \rightarrow 4 \text{H}_3\text{PO}_4$  Syn.
- ~~4~~  $\text{NH}_3 + \frac{7}{2} \text{O}_2 \rightarrow \frac{4}{2} \text{NO}_2 + \frac{6}{2} \text{H}_2\text{O}$  Comb.
- $\frac{1}{2} \text{C}_7\text{H}_8\text{O}_2 + \frac{15}{2} \text{O}_2 \rightarrow \frac{14}{2} \text{CO}_2 + \frac{6}{2} \text{H}_2\text{O}$  Comb.

# Answer Key To Classifying Chemical Reactions

**Michael Rajnik**



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