

Assignment #3 – Compound Names and Formulas
Single and Multi-valent ions

A. Write the correct formula for the following compounds.

- gallium oxide Ga_2O_3
- mercury (I) chloride $HgCl$
- tin (IV) bromide $SnBr_4$
- indium oxide In_2O_3
- cadmium sulphide CdS
- magnesium oxide MgO
- potassium nitride K_3N
- iron (III) sulphide Fe_2S_3
- gold (III) bromide $AuBr_3$
- zinc oxide ZnO

- silicon oxide SiO_2
- aluminum bromide $AlBr_3$
- tin (II) sulphide SnS
- germanium bromide $GeBr_4$
- nickel (II) oxide NiO
- strontium astatide $SrAt_2$
- aluminum oxide Al_2O_3
- copper (II) oxide CuO
- lawrencium fluoride LrF_3
- actinium oxide Ac_2O_3

B. Name the following compounds correctly. Only include the Roman Numeral in the name if the first element has 2 or more combining capacities.

- $CaBr_2$ Calcium Bromide
- K_2O Potassium Oxide
- Fe_2O_3 Iron (III) Oxide
- ZnF_2 Zinc Fluoride
- $CuCl$ Copper (I) Chloride
- Ni_2O_3 Nickel (III) Oxide
- $AuCl$ Gold (I) Chloride
- $HgBr$ Mercury (I) Bromide

- $AsCl_5$ Arsenic (V) Chloride
- Eu_2O_3 Europium (III) Oxide
- $NdBr_3$ Neodymium Bromide
- Cu_3N_2 Copper (II) Nitride
- CoS Cobalt (II) Sulphide
- ZrO_2 Zirconium Oxide
- Li_2O Lithium Oxide
- NaH Sodium Hydride

Answer Key To Covalent Compound Naming

Bill W. Tillery



Answer Key To Covalent Compound Naming:

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