

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

What happened to  
Avogadro when he  
got bit by  $6.02 \times 10^{23}$   
mosquitoes?  
He got Mol-aria

## POGIL- Stoichiometry

### How do chemists use balanced chemical equations?

#### Why?

Chemists use balanced chemical equations as a basis to calculate how much reactant is needed or product is formed in a reaction. This is called **Stoichiometry- (stoi-key-ah-meh-tree)** Another way of looking at it is using the mole ratio from the balanced equation and information about one compound in the reaction to determine information about another compound in the equation. A mole ratio is a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles. In chemical calculations, mole ratios are used to convert between moles one thing and moles of another.

**Purpose:** In this activity we will address the question: *How do I convert between different chemical species in a given reaction?*

### MODEL 1: Mole-to-Mole Stoichiometry (1-step)

#### Example: mole-to-mole

**Question:** Given the following equation:



How many moles of  $\text{H}_2$  are needed to produce 34.8 moles of  $\text{NH}_3$ ?

**Answer:**

$$\frac{34.8 \text{ mol NH}_3}{1} \times \frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3} = 52.2 \text{ mol H}_2$$

OR

$$34.8 \text{ mol NH}_3 \times \frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3} = 52.2 \text{ mol H}_2$$

#### Key Questions:

1. What was the given in the example above?
2. Where did the  $\frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3}$  come from?
3. Double check to see if 52.2 moles of  $\text{H}_2$  is the correct answer. SHOW the math below.

# Answer For Basic Stoichiometry Pogil Activity

**L Cohen**



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