

Stoichiometry Worksheet

I. ONE REACTANT IN EXCESS – COMBUSTION

You burn an excess of butane (C_4H_{10}) in a limited amount of oxygen (O_2). If 67.2 L of CO_2 gas is formed, what volume of oxygen gas was present at the start of the reaction?

STEP 1: Balance the reaction:



STEP 2: Use a molar conversion to convert quantity of product given to moles:

$$67.2 \text{ L } CO_2 * \frac{(1 \text{ mol } CO_2)}{(22.4 \text{ L } CO_2)} = 3 \text{ mol } CO_2$$

STEP 3: Use a molar ratio to calculate the moles of reactant that I'm asking about:

$$3 \text{ mol } CO_2 * \frac{(13 \text{ mol } O_2)}{(8 \text{ mol } CO_2)} = 4.875 \text{ mol } O_2$$

STEP 4: Use a molar conversion to convert moles of reactant to the units I ask for:

$$4.875 \text{ mol } O_2 \frac{(22.4 \text{ L } O_2)}{(1 \text{ mol } O_2)} = 109.2 \text{ L } O_2 = \boxed{109 \text{ L } O_2}$$

II. ONE REACTANT IN EXCESS – SINGLE DISPLACEMENT

14 g of lithium is combined with excess K_2SO_4 . What is the mass of the new ionic compound that is formed? How do you know that this reaction will take place?

STEP 1: Balance the reaction:



Reaction occurs because Li is higher up than K on the activity series.

STEP 2: Use a molar conversion to convert quantity of reactant given to moles:

$$14 \text{ g } Li * \frac{(1 \text{ mol } Li)}{(7 \text{ g } Li)} = 2 \text{ mol } Li$$

STEP 3: Use a molar ratio to calculate moles of product that I'm asking about:

$$2 \text{ mol } Li * \frac{(1 \text{ mol } Li_2SO_4)}{(2 \text{ mol } Li)} = 1 \text{ mol } Li_2SO_4$$

STEP 4: Use a molar conversion to convert moles of product to units that I ask for:

$$1 \text{ mol } Li_2SO_4 * \frac{(110 \text{ g } Li_2SO_4)}{(1 \text{ mol } Li_2SO_4)} = \boxed{110 \text{ g } Li_2SO_4}$$

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