

# Limiting Reactants Practice



a) What is the limiting reactant if 0.866 mol of NO is mixed with 0.503 mol of O<sub>2</sub> to create NO<sub>2</sub>?

$$0.866 \text{ mol NO} \times \frac{2 \text{ mol NO}_2}{2 \text{ mol NO}} = 0.866 \text{ mol NO}_2 \text{ (product made)}$$

limiting reactant

$$0.503 \text{ mol O}_2 \times \frac{2 \text{ mol NO}_2}{1 \text{ mol O}_2} = 1.006 \text{ mol NO}_2$$

b) Calculate the number of moles of NO<sub>2</sub> produced. (Answer should already be solved from the previous questions)

**0.866 mol NO<sub>2</sub>**



a) If 7.40 g of O<sub>3</sub> reacts with 0.670 g of NO, which compound will be the limiting reactant?

O<sub>3</sub>

$$O: 3 \times 15.999 = 47.997 \text{ g/mol}$$

$$\Rightarrow 1 \text{ mol O}_3 = 47.997 \text{ g O}_3$$

$$7.40 \text{ g O}_3 \times \frac{1 \text{ mol O}_3}{47.997 \text{ g O}_3} \times \frac{1 \text{ mol NO}_2}{1 \text{ mol O}_3} \times \frac{46.005 \text{ g NO}_2}{1 \text{ mol NO}_2} = 7.093 \text{ g NO}_2$$

NO

$$N: 1 \times 14.007 = 14.007$$

$$O: 1 \times 15.999 = 15.999$$

$$\Rightarrow 1 \text{ mol NO} = 30.006 \text{ g/mol}$$

$$0.670 \text{ g NO} \times \frac{1 \text{ mol NO}}{30.006 \text{ g NO}} \times \frac{1 \text{ mol NO}_2}{1 \text{ mol NO}} \times \frac{46.005 \text{ g NO}_2}{1 \text{ mol NO}_2} = 1.027 \text{ g NO}_2$$

limiting reactant

b) How many grams of NO<sub>2</sub> will be produced? (Answer should already be solved from the previous questions)

**1.027 g NO<sub>2</sub>**

NO<sub>2</sub>

$$N: 1 \times 14.007 = 14.007$$

$$O: 2 \times 15.999 = 31.998$$

$$\Rightarrow 1 \text{ mol NO}_2 = 46.005 \text{ g/mol}$$

c) Calculate mass, in grams, of the excess reactant remaining at the end of the reaction. How many moles of excess reactant would this be?

$$0.670 \text{ g NO} \times \frac{1 \text{ mol NO}}{30.006 \text{ g NO}} \times \frac{1 \text{ mol O}_3}{1 \text{ mol NO}} \times \frac{47.997 \text{ g O}_3}{1 \text{ mol O}_3} = 1.072 \text{ g O}_3 \text{ used!}$$

$$\frac{7.40 \text{ g O}_3}{\text{original}} - \frac{1.072 \text{ g O}_3}{\text{used}} = 6.328 \text{ g O}_3 \text{ leftover}$$

# Answer Key To Pogil 33 Limiting Reactants

**Michal Rosen-Zvi**



**Answer Key To Pogil 33 Limiting Reactants:**

## Whispering the Strategies of Language: An Mental Quest through **Answer Key To Pogil 33 Limiting Reactants**

In a digitally-driven world where displays reign supreme and immediate conversation drowns out the subtleties of language, the profound strategies and mental subtleties hidden within phrases frequently get unheard. However, located within the pages of **Answer Key To Pogil 33 Limiting Reactants** a fascinating fictional value pulsing with fresh thoughts, lies an exceptional journey waiting to be undertaken. Published by a talented wordsmith, that charming opus encourages readers on an introspective trip, softly unraveling the veiled truths and profound influence resonating within the very material of each and every word. Within the emotional depths of the moving review, we can embark upon a heartfelt exploration of the book is key subjects, dissect its charming publishing style, and succumb to the effective resonance it evokes heavy within the recesses of readers hearts.

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