

FIRST SEMESTER CHEMISTRY FINAL EXAMINATION FOR GRADE 12TH**CHOOSE THE CORRECT ANSWERS**

- If the solute-solvent interactions are greater than the solute-solute and solvent-solvent interactions, what will be the total vapor pressure of the solution?
 - greater than that calculated from Raoult's law
 - less than that calculated from Raoult's law
 - The same as calculated from Raoult's law
 - Raoult's law cannot be applied for such interaction
- What is the freezing point of a solution of 250g of CaCl_2 in 1.0kg of water (K_f for water = $-1.86^\circ\text{C}/\text{kg}/\text{mol}$)
 - -1.3°C
 - -6.5°C
 - -9.0°C
 - -12.55°C
- What is the molarity of a 5 g of hydrogen peroxide in 100mL solution that is used for hair bleaching
 - 0.015M
 - 0.15M
 - 1.5M
 - 3M
- What is the osmotic pressure, in atm, of a 0.30M solution of glucose in water that is used for intravenous infusion at body temperature, 37°C ?
 - 0.911 atm
 - 773.2atm
 - 7.6atm
 - 76atm
- Given the unbalanced equation
 $\text{KMnO}_4 + \text{KI} + \text{H}_2\text{SO}_4 \rightarrow \text{K}_2\text{SO}_4 + \text{MnSO}_4 + \text{I}_2 + \text{H}_2\text{O}$ How many grams of KMnO_4 are needed to make 500mL of 0.250N solution?
 - 3.95g
 - 1.49g
 - 1.58g
 - 63.2g
- In which of the following cases will the dissolution of sugar be the most rapid?
 - Sugar crystals in hot water
 - Sugar crystals in cold water
 - Powdered sugar in hot water
 - Powdered sugar in cold water
- Which of the following is not a colligative property
 - Osmotic pressure
 - Vapour pressure
 - Boiling point
 - Freezing point depression
- What is the normality of a 1M solution of H_3PO_4 given the following reaction
 $\text{H}_3\text{PO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{HPO}_4 + \text{H}_2\text{O}$
 - 3N
 - 2N
 - 6N
 - 1N
- A given solution contains 100g of NaCl and 900g of water, what is the mole fraction of NaCl
 - 0.0331
 - 0.9669
 - 0.3
 - 0.7
- A 0.029m aqueous solution of K_2SO_4 has a freezing point of -0.14°C , what are the value of i factors ($K_{f\text{H}_2\text{O}} = -1.86^\circ\text{C}/\text{m}$)
 - 2
 - 4
 - 2.3
 - 2.6
- Which of the following combination never form a buffer solution
 - HF/F^-
 - $\text{HCOOH}/\text{HCOO}^-$
 - $\text{HCO}_3^-/\text{CO}_3^{2-}$
 - $\text{HNO}_3/\text{NO}_3^-$
- Which one of the following are true about the graph of weak acid and strong base
 - The PH is equal to 7 at equivalence point
 - The graph indicates the titration of strong base with strong acid
 - The PH is greater than 7 at equivalence point
 - at half neutralization $\text{PH} = \text{PKa}$
- A solution in which propionic acid is 0.94% ionized has a PH of 2.85, what is the values of the acid ionization constant (K_a) for propionic acid
 - 4.2×10^{-5}
 - 6.8×10^{-4}
 - 1.3×10^{-5}
 - 8.7×10^{-7}
- The acid-base indicator bromocresol green is a weak acid. The yellow acid and blue base forms of the indicator are present in equal concentrations in a solution when the PH is 4.68, what is the PKa of bromocresol green
 - 4.48
 - 4.68
 - 5.68
 - 6.68
- For the step wise dissociation of aqueous H_3PO_4 , which of the following is not conjugate acid-base pair
 - HPO_4^{2-} and H_2PO_4^-
 - H_3PO_4 and H_2PO_4^-
 - H_2PO_4^- and HPO_4^{2-}
 - H_2PO_4^- and PO_4^{3-}
- To 0.2M solution of a weak monoprotic acid HA, enough quantity of its sodium salt NaA, was dissolved to give a concentration of 0.2M of the salt, what will be the acid concentration $[\text{H}_3\text{O}^+]$ in the final solution (K_a of HA = 1.8×10^{-5})
 - $1.8 \times 10^{-5}\text{M}$
 - $1.8 \times 10^{-4}\text{M}$
 - 2.8×10^{-5}
 - 2.8×10^{-4}
- Consider the following equilibria
 - $\text{HCO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3 + \text{OH}^-$
 - $\text{NH}_4^+ + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{NH}_3$
 - $\text{HSO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{O} + \text{H}_2\text{SO}_3$
 In which of the equilibrium does water acts as a Bronsted-lowry base?
 - I only
 - II only
 - II and III
 - I, II, III
- To a 50ml of a buffer solution that is 0.25M in HF and 0.1M in NaF, if 5ml of 0.05M NaOH is added, what is the final PH of this buffer solution, $\text{PKa HF} = 3.17$
 - 2.77
 - 2.80
 - 2.74
 - 3.17
- When sodium acetate is added to an aqueous solution of acetic acid its PH
 - increases
 - Decreases
 - Remain constant
 - may increases or decreases
- How is a buffer solution prepared
 - By mixing a weak acid and strong acid
 - By mixing a weak acid and its conjugate base
 - By mixing a strong acid and its conjugate base
 - By mixing strong base and its conjugate acid
- Which one of the followings are not intensive properties
 - color
 - length
 - density
 - melting point

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**Thirumalaisamy P.
Velavancorresponding**

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