

Chapter 7 The Mole and Chemical Composition

Section 1 Avogadro's Number and Molar Conversions

Objectives

- Identify the mole as the unit used to count particles, whether atoms, ions, or molecules.
- Use Avogadro's number to convert between amount in moles and number of particles.
- Solve problems converting between mass, amount in moles, and number of particles using Avogadro's number and molar mass.

Avogadro's Number and the Mole

- The SI unit for amount is called the mole (mol). A **mole** is the number of atoms in exactly 12 grams of carbon-12.
- Scientists use the mole to make counting large numbers of particles easier.
- The number of particles in a mole is called **Avogadro's Number**. Avogadro's number is $6.02214199 \times 10^{23}$ units/mole.
- The mole is used to count out a given number of particles, whether they are atoms, molecules, formula units, ions, or electrons.
- The mole is just one kind of counting unit:
 - 1 dozen = 12 objects
 - 1 hour = 3600 seconds
 - 1 mole = 6.022×10^{23} particles

Amount in Moles Can Be Converted to Number of Particles

- Counting units are used to make conversion factors.
- The definition of one mole is

$$6.022 \times 10^{23} \text{ particles} = 1 \text{ mol}$$

- The conversion factor is

$$\frac{6.022 \times 10^{23} \text{ particles}}{1 \text{ mol}} = 1$$

Choose the Conversion Factor That Cancels the Given Units

- All conversion factors are equal to 1, so you can use them to convert among different units.
- You can tell which conversion factor to use, because the needed conversion factor should cancel the units of the given quantity to give you the units of the answer or the unknown quantity.



Converting Amount in Moles to Number of Particles

Sample Problem A

Find the number of molecules in 2.5 mol of sulfur dioxide.

Sample Problem A Solution (Page Number 228)

$$2.5 \text{ mol SO}_2 \times ? = ? \text{ molecules SO}_2$$

You are converting from the unit *mol* to the unit *molecules*. The conversion factor must have the units of *molecules/mol*.

You use 6.022×10^{23} molecules/1 mol.

$$2.5 \text{ mol SO}_2 \times ? = ? \text{ molecules SO}_2$$

Number of Particles Can Be Converted to Amount in Moles

- The reverse calculation is similar to that in Sample Problem A but the conversion factor is inverted to get the correct units in the answer.
 - example: How many moles are 2.54×10^{22} iron(III) ions?

Chapter 7 The Mole And Chemical Composition

Sidney William Benson



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