

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ **lattice energy** _____ a. ionic bond
 _____ **bond energy** _____ b. covalent bond
 _____ **enthalpy of vaporization** _____ c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ **ionic** _____ a. H—F _____ **nonpolar covalent** _____ d. H—H
 _____ **ionic** _____ b. Na—Cl _____ **polar covalent** _____ e. H—C
 _____ **polar covalent** _____ c. H—O _____ **polar covalent** _____ f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond. Increased polarity in H₂O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

Chapter 8 2 Workbook Covalent Bonding Answer Key

Jessica J Manson



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Molecular Biology of the Cell 6E - The Problems Book John Wilson, Tim Hunt, 2014-11-21 The Problems Book helps students appreciate the ways in which experiments and simple calculations can lead to an understanding of how cells work by introducing the experimental foundation of cell and molecular biology Each chapter reviews key terms tests for understanding basic concepts and poses research based problems The Problems Book has been

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