

Chapter 8 ACID-BASE CHEMISTRY & EQUILIBRIUM

Zumdahl ch. 14, 15.1-15.5

All states are aqueous unless indicated otherwise.

I. STRUCTURE DETERMINES FUNCTION – the notes below are covered in the Acid/Base Inquiry.

ACID STRUCTURE & STRENGTH		
TREND	DIAGRAM	EXPLANATION
<p>BINARY ACIDS: INCREASE in strength L→R across a period. Binary acids INCREASE in strength DOWN a group.</p>		<p>The acid strength is related to the bond strength (down a group) as well as the electronegativity of the element bonded to the "H" (across a period)</p>
ACID STRUCTURE & STRENGTH		
TREND	DIAGRAM	EXPLANATION
<p>OXYACIDS: INCREASE in strength as the number of oxygen atoms on the central atoms INCREASES</p>		<ul style="list-style-type: none"> ✓ Nitric acid has one more oxygen. ✓ More electron density is pulled from the O-H bond. ✓ H⁺ leaves more readily ✓ HNO₃ is thus a stronger acid.
<p>OXYACID: BE CAREFUL! Not all "H" atoms are acidic!</p>	<p>Phosphoric acid, H₃PO₄ $K_{a1} = 7.5 \times 10^{-3}$ $K_{a2} = 6.2 \times 10^{-8}$ $K_{a3} = 4.8 \times 10^{-13}$</p> <p>Phosphorous acid, H₃PO₃ $K_{a1} = 2.5 \times 10^{-2}$ $K_{a2} = 2.4 \times 10^{-7}$</p>	<p>The element bonded directly to the "H" must be sufficiently electronegative to weaken the bond to the "H". A bond between P-H is not acidic. Therefore H₃PO₄ is TRI-protic and H₃PO₃ is DI-protic</p>

Chapter 8 Solution Acids And Bases Wordwise

Rachel S Tattersall



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