

# Naming Ionic Compounds

What are the structural units that make up ionic compounds and how are they named?

## Why?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as NaCl. In talking about chemistry however, it is a bit tacky to say "en-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

## Model 1 - Ion Charges for Selected Elements

|   |                 |                  |                                      |                                      |                                     |   |                  |                                      |                 |                 |                 |  |  |
|---|-----------------|------------------|--------------------------------------|--------------------------------------|-------------------------------------|---|------------------|--------------------------------------|-----------------|-----------------|-----------------|--|--|
| 1 | H <sup>+</sup>  |                  |                                      |                                      |                                     |   |                  |                                      |                 |                 |                 |  |  |
| 2 | Li <sup>+</sup> | Be <sup>2+</sup> |                                      |                                      |                                     |   |                  |                                      | N <sup>3-</sup> | O <sup>2-</sup> | F <sup>-</sup>  |  |  |
| 3 | Na <sup>+</sup> | Mg <sup>2+</sup> | Transition elements                  |                                      |                                     |   | Al <sup>3+</sup> |                                      | P <sup>3-</sup> | S <sup>2-</sup> | Cl <sup>-</sup> |  |  |
| 4 | K <sup>+</sup>  | Ca <sup>2+</sup> | Fe <sup>2+</sup><br>Fe <sup>3+</sup> | Ni <sup>2+</sup><br>Ni <sup>3+</sup> | Cu <sup>+</sup><br>Cu <sup>2+</sup> | Zn <sup>2+</sup>                                  |                  |                                      |                 |                 | Br <sup>-</sup> |  |  |
| 5 | Rb <sup>+</sup> | Sr <sup>2+</sup> |                                      |                                      | Ag <sup>1+</sup>                    |   |                  | Sn <sup>2+</sup><br>Sn <sup>4+</sup> |                 |                 | I <sup>-</sup>  |  |  |
| 6 |                 | Ba <sup>2+</sup> |                                      |                                      |                                     | Hg <sub>2</sub> <sup>2+</sup><br>Hg <sup>2+</sup> |                  | Pb <sup>2+</sup><br>Pb <sup>4+</sup> |                 |                 |                 |  |  |

← Cations →
 
 ← Anions →

1. Consider the ions of potassium (K) and sulfur (S). Write the chemical formulas for the ionic compound involving these ions, using the simplest ratio(s) of potassium (K) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.



2. Consider the ions of iron (Fe) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of iron (Fe) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.



## Model 2 - Ionic Compound Names (Metals that form one ion)

NaCl Sodium chloride

Zn<sub>3</sub>P<sub>2</sub> Zinc phosphide

CaS Calcium sulfide

Al<sub>2</sub>O<sub>3</sub> Aluminum oxide

Ag<sub>2</sub>S Silver sulfide

SrCl<sub>2</sub> Strontium chloride

3. Circle the symbol for the metal in each of the compounds in Model 2. If doing electronically, you can just **BOLD** the symbol. Remember your PT!

# Chapter 9 Chemistry Naming Formulas Answer Key

**Kimberley Waldron**



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