

9.1 Assessment

- The positive charge of a group A metal is equal to its group number. The charge of a group A nonmetal is the group number minus 8.
- from the number of electrons lost
- ite or -ate
- Group 1A metals, 1+; Group 3A (aluminum), 3+; Group 5A nonmetals, 3-
- A monoatomic anion is a single atom with a negative charge; a polyatomic anion is two or more bound atoms with a negative charge.
- a.) K^+ , cation, potassium ion
b.) O^{2-} anion, oxide ion
c.) Sn^{2+} cation, tin (II)
d.) Br^- anion, bromide ion
e.) Be^{2+} cation beryllium ion
f.) Co^{3+} cation, cobalt(III) ion
- a.) NH_4^+
b.) Sn^{2+}
c.) CrO_4^{2-}
d.) NO_3^-

9.2 Assessment

- Write the name of the cation followed by the name of the anion.
- Write the symbol of the cation followed by the symbol of the anion, then use subscripts to balance the charges.
- Write the symbol for the metal ion followed by the formula of the polyatomic ion and balance the charges. Name the cation first followed by the anion.
- a.) $BeCl_2$
b.) Cs_2S
c.) NaI
d.) SrO
- a.) $Cr(NO_3)_3$
b.) $NaClO_4$
c.) $Mg(HCO_3)_2$
d.) $Ca(C_2H_3O_2)_2$
- a.) incorrect; charges are not balanced $MgSO_4$
b.)
c.) incorrect; charges are not balanced $BeCl_2$

9.3 Assessment

- Prefixes indicate the number of atoms of each element in a molecule of the compound.

Chapter Assessment Chemistry Answers

Kaplan, Michael Palmer



Chapter Assessment Chemistry Answers:

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Table of Contents Chapter Assessment Chemistry Answers

1. Understanding the eBook Chapter Assessment Chemistry Answers
 - The Rise of Digital Reading Chapter Assessment Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Assessment Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter Assessment Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter Assessment Chemistry Answers
 - Personalized Recommendations
 - Chapter Assessment Chemistry Answers User Reviews and Ratings

- Chapter Assessment Chemistry Answers and Bestseller Lists
- 5. Accessing Chapter Assessment Chemistry Answers Free and Paid eBooks
 - Chapter Assessment Chemistry Answers Public Domain eBooks
 - Chapter Assessment Chemistry Answers eBook Subscription Services
 - Chapter Assessment Chemistry Answers Budget-Friendly Options
- 6. Navigating Chapter Assessment Chemistry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter Assessment Chemistry Answers Compatibility with Devices
 - Chapter Assessment Chemistry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter Assessment Chemistry Answers
 - Highlighting and Note-Taking Chapter Assessment Chemistry Answers
 - Interactive Elements Chapter Assessment Chemistry Answers
- 8. Staying Engaged with Chapter Assessment Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter Assessment Chemistry Answers
- 9. Balancing eBooks and Physical Books Chapter Assessment Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter Assessment Chemistry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter Assessment Chemistry Answers
 - Setting Reading Goals Chapter Assessment Chemistry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter Assessment Chemistry Answers
 - Fact-Checking eBook Content of Chapter Assessment Chemistry Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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