

The Structure of the Atom

Section 4.1 Early Ideas About Matter

pages 102–105

Section 4.1 Assessment

page 105

1. Contrast the methods used by the Greek philosophers and Dalton to study the atom.

Greek philosophers could not conduct experiments to verify their hypothesis, whereas Dalton could make careful measurements.

2. Define *atom* using your own words.

An atom is the smallest component of an element that exhibits all of the characteristic properties of that element.

3. **Summarize** Dalton's atomic theory.

Matter is composed of small particles called atoms that are indivisible. Atoms of an element are identical in size, mass, and chemical properties. Atoms of a specific element are different from atoms of another element. Different atoms combine in simple whole-number ratios to form compound. In a chemical reaction, atoms are separated, combined, or rearranged.

4. Explain how Dalton's theory of the atom and the conservation of mass are related.

Dalton explained that atoms are not created or destroyed in chemical reactions, but only rearranged.

5. **Apply** Six atoms of Element A combine with 8 atoms of Element B to produce six compound particles. How many atoms of Elements A and B does each particle contain? Are all atoms used to form compounds?

Each compound contains 1 atom A and 1 atom B. 2 atoms of element B are not used.

6. Design a concept map that compares and contrasts the atomic theories proposed by Democritus and John Dalton.

Concept maps will vary, but should reflect the following summary. Both believed: matter composed of extremely small particles called atoms; all atoms of a given element are identical, but differ from the atoms of other elements; atoms could not be created, divided, or destroyed; apparent changes in matter result from changes in the groupings of atoms. Democritus further believed that matter is composed of empty space through which atoms move, that different kinds of atoms come in different sizes and shapes, and that the differing properties of atoms are due to the size, shape, and movement of the atoms. Dalton further specified that different atoms combine in simple whole number ratios to form compounds.

Section 4.2 Defining the Atom

pages 106–114

Section 4.2 Assessment

page 114

7. Describe the structure of a typical atom. Identify where each subatomic particle is located.

A typical atom consists of a central, small, dense nucleus containing protons and neutrons. The nucleus is surrounded by a cloud of negatively charged electrons.

8. **Compare and contrast** Thomson's plum pudding atomic model with Rutherford's nuclear atomic model.

Thomson's plum pudding model describes atoms as spherical particles with uniformly distributed positive charge in which individual, negatively charged, electrons are located in fixed positions. In contrast, Rutherford's model states that an atom is mostly empty space, with a small, dense, central nucleus containing all of an atom's positive charge and most of its mass. The negatively charged electrons move through the empty space and are held in the atom by their attraction to the positively charged nucleus.

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