

Chapter 7: Chemical Quantities

Learning Objectives:

- 1) Use atomic mass from periodic table to calculate molecular mass of a molecule
- 2) Understand the relationship between the "amu" and the "mole"
- 3) Be able to convert between grams and moles using molecular mass
- 4) Be able to convert between moles and # particles using Avogadro's number
- 5) Differentiate between empirical and molecular formula
- 6) Be able to determine the empirical formula given experimentally determined mass percentage values for a given substance
- 7) Be able to compute the molecular formula given objective (6) and the experimentally determined molecular mass for a given substance

Molecular Mass

- Atoms have an atomic mass. Shouldn't a molecule have a molecular mass? Of course. Determining the molecular mass is as easy as it should be.

- H has an atomic mass of 1.00794 amu
- H₂ has a molecular mass of 2(1.00794amu) = 2.01588 amu

- NaCl has a MM of
$$\begin{array}{r} 22.99 \text{ amu} \\ + 35.45 \text{ amu} \\ \hline 58.44 \text{ amu} \end{array}$$

- Calculate the molecular mass for the following

a) Sulfur dioxide

b) ammonia

c) Calcium chlorate

| | |
|----|------------|
| C | 12.01 amu |
| N | 14.01 amu |
| O | 16.00 amu |
| F | 19.00 amu |
| S | 32.065 amu |
| Cl | 35.45 amu |
| Na | 22.99 amu |
| Ca | 40.08 amu |

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