

**9.1 Assessment**

- The positive charge of a group A metal is equal to its group number. The charge of a group A nonmetal is the group number minus 8.
- from the number of electrons lost
- ite or -ate
- Group 1A metals, 1+; Group 3A (aluminum), 3+; Group 5A nonmetals, 3-
- A monoatomic anion is a single atom with a negative charge; a polyatomic anion is two or more bound atoms with a negative charge.
- a.)  $K^+$ , cation, potassium ion  
b.)  $O^{2-}$  anion, oxide ion  
c.)  $Sn^{2+}$  cation, tin (II)  
d.)  $Br^-$  anion, bromide ion  
e.)  $Be^{2+}$  cation beryllium ion  
f.)  $Co^{3+}$  cation, cobalt(III) ion
- a.)  $NH_4^+$   
b.)  $Sn^{2+}$   
c.)  $CrO_4^{2-}$   
d.)  $NO_3^-$

**9.2 Assessment**

- Write the name of the cation followed by the name of the anion.
- Write the symbol of the cation followed by the symbol of the anion, then use subscripts to balance the charges.
- Write the symbol for the metal ion followed by the formula of the polyatomic ion and balance the charges. Name the cation first followed by the anion.
- a.)  $BeCl_2$   
b.)  $Cs_2S$   
c.)  $NaI$   
d.)  $SrO$
- a.)  $Cr(NO_3)_3$   
b.)  $NaClO_4$   
c.)  $Mg(HCO_3)_2$   
d.)  $Ca(C_2H_3O_2)_2$
- a.) incorrect; charges are not balanced  $MgSO_4$   
b.)  
c.) incorrect; charges are not balanced  $BeCl_2$

**9.3 Assessment**

- Prefixes indicate the number of atoms of each element in a molecule of the compound.

# Chapter Chemistry Assessment Answers

**Cynthia Knable**



## **Chapter Chemistry Assessment Answers:**

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