

# Chapter 10 Chemical Quantities Review

## Part 1: Vocabulary

Label the structure corresponding to the term in the parentheses.

- 1.** mole: gram-molecular mass
- 2.** empirical formula: simplest formula
- 3.** molar mass: molecular mass and gram-molecular mass
- 4.**  $10^{-24}$ : unit of mass
- 5.** a description of the atoms present in a substance is a composition.
- 6.** the fixed whole number ratio of the atoms of the elements in a compound is a simplest formula.
- 7.** molecular structure is the arrangement of atoms in a molecule.
- 8.** the smallest practical amount of a compound is a mole.
- 9.** the fixed composition of a substance is a simplest formula.
- 10.**  $10^{-24}$  particles is a unit of mass.
- 11.** the smallest amount of a substance is a mole.
- 12.** a description of the atoms present in a substance is a composition.

## Part 2: Section Concept Review

### 10.1 The Mole: A Fundamental

**Completion** Use the completion method to check your knowledge of the terms and facts concerning the concepts involved in the section. Each blank space represents a term. Don't guess; or write.

Knowing what mole is, counting, weighing, or measuring a volume to a definite quantity called the mole. The number of particles in one mole of a substance is Avogadro's number. It is the ratio of mass of a compound, molecule and together the molar mass of the substance of the compound. When you substitute the unit given for mass you obtain the molar mass of the compound. These are fixed values represented particles are made of the substance.

#### Multiple Choice

- 10.1** a. smallest practical amount of a compound: mole
- 10.2** a. description of the atoms present in a substance: composition
- 10.3** a. The smallest practical amount of a substance is the mole.
- 10.4** a. the fixed composition of a substance is a simplest formula.
- 10.5** a. The mole mass of nitrogen gas is 28 g.

#### True/False

# Chapter Chemical Quantities Guided Reading Answer Key

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