

AP Chemistry: Grade 11. Honors/Preparatory

Unit 12: Electrochemistry

Activity 12: Electrochemistry: Voltaic Cells and Batteries



Activity 12: Electrochemistry: Voltaic Cells and Batteries

Objectives: Develop a model that allows for prediction of the direction of spontaneous redox reactions.

Prerequisites:

- 1. Use the periodic table, knowledge of electronegativity and bond dissociation energy to predict reactivity.
- 2. Use the periodic table, knowledge of electronegativity and bond dissociation energy to predict reactivity.
- 3. Use the periodic table, knowledge of electronegativity and bond dissociation energy to predict reactivity.
- 4. Use the periodic table, knowledge of electronegativity and bond dissociation energy to predict reactivity.
- 5. Use the periodic table, knowledge of electronegativity and bond dissociation energy to predict reactivity.
- 6. Use the periodic table, knowledge of electronegativity and bond dissociation energy to predict reactivity.

Materials:

Activity 1: Develop the electrochemical cell, and the sign of the cell potential is related to the direction of the spontaneous redox reaction.

Activity 2:

The diagram shows four electrochemical cells, each with two half-cells connected by a salt bridge. The half-cells are labeled with their respective metals and ions:

- (a) Zinc half-cell (Zn | Zn²⁺) and Silver half-cell (Ag⁺ | Ag)
- (b) Zinc half-cell (Zn | Zn²⁺) and Copper half-cell (Cu²⁺ | Cu)
- (c) Zinc half-cell (Zn | Zn²⁺) and Silver half-cell (Ag⁺ | Ag)
- (d) Zinc half-cell (Zn | Zn²⁺) and Silver half-cell (Ag⁺ | Ag)

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