

Unit 8 Review

1. The balanced equation  $\text{P}_4(\text{s}) + 6\text{H}_2(\text{g}) \rightarrow 4\text{PH}_3(\text{g})$  tells us that 4.0 mol  $\text{H}_2$

A) reacts with 2.0 mol  $\text{P}_4$ .  
 B) produces 8.0 mol  $\text{PH}_3$ .  
 C) cannot react with phosphorus.  
 D) produces 2.7 mol  $\text{PH}_3$ .  
 E) reacts with 4.0 mol  $\text{P}_4$ .

2. A 1.0-mol sample of  $\text{Fe}_2\text{O}_3$  was decomposed according to the equation:

$$2\text{Fe}_2\text{O}_3(\text{s}) \rightarrow 2\text{Fe}(\text{s}) + 3\text{O}_2(\text{g})$$

How many moles of  $\text{O}_2$  are formed assuming 100% yield?

- A) 0.0 mol  
 B) 1.5 mol  
 C) 1.7 mol  
 D) 0.7 mol  
 E) 2.0 mol

3. The rusting of iron is represented by the equation  $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$ . If you have a 1.00-mol sample of iron, how many moles of  $\text{Fe}_2\text{O}_3$  will there be after the iron has rusted completely?

- A) 0.527 mol  
 B) 0.750 mol  
 C) 1.00 mol  
 D) 1.50 mol  
 E) 2.27 mol

4. Consider the following reaction, where X represents an unknown element:

$$8\text{X}_2(\text{s}) + 2\text{B}_2\text{O}_3(\text{s}) \rightarrow \text{B}_2\text{X}_6(\text{s}) + 3\text{O}_2(\text{g})$$

If 175 g of X reacts completely with diboron trioxide to produce 1.43 mol of  $\text{B}_2\text{X}_6$ , what is the identity of X?

- A) Ge  
 B) Mg  
 C) Si  
 D) Sn  
 E) C

5. Refer to the following equation:

$$4\text{NH}_3(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$$

How many molecules of  $\text{NO}$  are produced when 3.02 mol of ammonia is completely reacted?

- A) 12.08  
 B)  $3.84 \times 10^{23}$   
 C)  $1.82 \times 10^{23}$   
 D) 1.39  
 E) none of these

6. In the reaction  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ , how many moles of ammonia would be produced from 1.00 mol of hydrogen and excess nitrogen?

- A) 1.42 mol  
 B) 3.27 mol  
 C) 0.727 mol  
 D) 2.18 mol  
 E) 0.242 mol

7. Refer to the following unbalanced equation:

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

What mass of oxygen ( $\text{O}_2$ ) is required to react completely with 13.1 g of  $\text{C}_2\text{H}_6$ ?

- A)  $2.98 \times 10^2$  g  
 B) 23.1 g  
 C) 4.86 g  
 D) 49.2 g  
 E) 0.192 g

8. How many molecules of carbon dioxide will be formed if 7.07 g of propane is burned in the following reaction?



- A)  $1.37 \times 10^{23}$  molecules  
 B)  $1.00 \times 10^{23}$  molecules  
 C)  $4.14 \times 10^{23}$  molecules  
 D)  $3.10 \times 10^{23}$  molecules  
 E)  $5.17 \times 10^{23}$  molecules

9. Calculate the mass of carbon dioxide produced from 15.7 g of octane,  $\text{C}_8\text{H}_{18}$ , in the following reaction:



- A) 79.2 g  $\text{CO}_2$   
 B) 158 g  $\text{CO}_2$   
 C) 89.1 g  $\text{CO}_2$   
 D) 1.258 g  $\text{CO}_2$   
 E) 1.400 g  $\text{CO}_2$

10. Calculate the molecules of oxygen required to react with 18.2 g of sulfur in the following reaction:



- A)  $1.03 \times 10^{23}$  molecules  $\text{O}_2$   
 B)  $2.28 \times 10^{23}$  molecules  $\text{O}_2$   
 C)  $1.71 \times 10^{23}$  molecules  $\text{O}_2$   
 D)  $3.42 \times 10^{23}$  molecules  $\text{O}_2$   
 E)  $5.13 \times 10^{23}$  molecules  $\text{O}_2$

# Multiple Choice Stoichiometry Test

**Paul Foglino**



**Multiple Choice Stoichiometry Test:**

## Decoding **Multiple Choice Stoichiometry Test**: Revealing the Captivating Potential of Verbal Expression

In a period characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its capability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Multiple Choice Stoichiometry Test**," a mesmerizing literary creation penned with a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring impact on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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