

Ch. 6 Section 2 (Describing Chemical Reactions)

Science

Vocabulary:

1. Chemical Equation	A short, easy way to show a chemical reaction, using symbols.
2. Conservation of Matter	The principle stating that matter is not created nor destroyed during a chemical reaction.
3. Open System	A system in which matter can enter from or escape to the surroundings.
4. Closed System	A system in which no matter is allowed to enter or leave.
5. Coefficient	A number in front of a chemical formula in an equation that indicates how many molecules or atoms of each reactant and product are involved in a reaction.
6. Synthesis	A chemical reaction in which two or more simple substances combine to form a new, more complex substance.
7. Decomposition	A chemical reaction that breaks down compounds into similar products.
8. Replacement	A reaction in which one element replaces another in a compound or when two elements in different compounds trade places.

9.

10. The reactants are those on the left side of the equation. The products are those on the right side of the equation. The plus signs are for adding things together, and the arrow is for the reaction to happen. The arrow goes from the reactants to the products.

11.

Reactants	Products	Products
Hydrogen gas and oxygen gas	Water	Hydrogen gas and oxygen gas
Hydrogen gas	Water	Hydrogen gas and oxygen gas

12.

13.

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SJ Ball



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Decoding **Chapter 1guided Notes Name 1aca Aeoe Describing Chemical**: Revealing the Captivating Potential of Verbal Expression

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