

**Focus 1**

**The properties of gases**

Answers to exercises

**Topic 1A The perfect gas**

- 1A.1** (a) 810 Torr  
 (b) 0.962 atm  
 (c) 0.222 atm  
 (d)  $1.03 \times 10^5$  Pa
- 1A.2** 92 kPa
- 1A.3** 2.25 kPa
- 1A.4** 4.33 mmol
- 1A.5** 660 bar
- 1A.6** 10.0 atm
- 1A.7** 418 kPa
- 1A.8** 173 kPa
- 1A.9**  $-243.6$  °C
- 1A.10** 394 K
- 1A.11** (a)  $3.6$  m<sup>3</sup>  
 (b)  $178$  m<sup>3</sup>
- 1A.12**  $3.4 \times 10^5$  m<sup>3</sup>
- 1A.13** benzene 0.355, toluene 0.645
- 1A.14** (a)  $1.32 \times 10^{-3}$  m<sup>3</sup>  
 (b)  $61.2 \times 10^3$  Pa

**Topic 1B The kinetic model of gases**

- 1B.1** (a)  $493$  m s<sup>-1</sup>  
 (b)  $615$  m s<sup>-1</sup>
- 1B.2** (a)  $v_{\text{mean}}(79 \text{ K}) = 647$  m s<sup>-1</sup>  
 $v_{\text{mean}}(315 \text{ K}) = 1.29$  km s<sup>-1</sup>  
 $v_{\text{mean}}(1500 \text{ K}) = 2.82$  km s<sup>-1</sup>  
 (b)  $v_{\text{mean}}(79 \text{ K}) = 323$  m s<sup>-1</sup>  
 $v_{\text{mean}}(315 \text{ K}) = 645$  m s<sup>-1</sup>  
 $v_{\text{mean}}(1500 \text{ K}) = 1.41$  km s<sup>-1</sup>
- 1B.3** 3.72
- 1B.4**  $m_{\text{N}_2}$  0.80 g  
 $m_{\text{He}}$  0.30 g
- 1B.5** 65 mPa

- 1B.6**  $\lambda = 0.97$  μm  
 $N = 120 \times 10^3$
- 1B.7** (a)  $7.0 \times 10^{10}$   
 (b)  $7.0 \times 10^8$   
 (c)  $7.0 \times 10^4$
- 1B.9** (a) 6.8 nm  
 (b) 68 nm  
 (c) 6.8 mm

**Topic 1C Real gases**

- 1C.1**  $z_c = 0.242$
- 1C.2** CH<sub>4</sub>  $z_c = 0.288$   
 C<sub>2</sub>H<sub>6</sub>  $z_c = 0.274$   
 C<sub>3</sub>H<sub>8</sub>  $z_c = 0.276$
- 1C.3** 2.15 MPa
- 1C.4**  $-3.37 \times 10^{-5}$  m<sup>3</sup> mol<sup>-1</sup>
- 1C.5**  $-3.49 \times 10^{-8}$  Pa<sup>-1</sup>
- 1C.6** (a) (i) 101 kPa  
 (ii) 83.1 MPa  
 (b) (i) 101 kPa  
 (ii) 206 MPa
- 1C.7**  $-0.608$  MPa
- 1C.8**  $b = 4.6 \times 10^{-5}$  m<sup>3</sup> mol<sup>-1</sup>  
 $Z = 0.66$
- 1C.9** 1842 K

Answers to problems

- 1.1** 0.50 m<sup>3</sup>
- 1.2** 28.82 g mol<sup>-1</sup>  
 130 kg
- 1.3** 6.8 kPa
- 1.4** 713 Torr
- 1.5**  $x_{\text{N}_2} = 0.7809$   
 $x_{\text{O}_2} = 0.2094$   
 $x_{\text{Ar}} = 0.0098$
- 1.6** 132 g mol<sup>-1</sup>
- 1.7** 16.4 g mol<sup>-1</sup>

# Chapter 1section The Properties Of Gases Answers

**RC Schank**



**Chapter 1 section The Properties Of Gases Answers:**

## Embracing the Track of Term: An Emotional Symphony within **Chapter 1section The Properties Of Gases Answers**

In some sort of consumed by screens and the ceaseless chatter of fast conversation, the melodic beauty and psychological symphony developed by the prepared term often fade into the backdrop, eclipsed by the relentless sound and disruptions that permeate our lives. Nevertheless, nestled within the pages of **Chapter 1section The Properties Of Gases Answers** a marvelous literary treasure filled with fresh emotions, lies an immersive symphony waiting to be embraced. Crafted by a masterful composer of language, this captivating masterpiece conducts readers on an emotional trip, well unraveling the concealed tunes and profound affect resonating within each carefully constructed phrase. Within the depths with this touching analysis, we shall examine the book is key harmonies, analyze its enthralling publishing model, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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