

Chapter 1: Solutions to Review Problems

Section 1.1

- 1) If the number of moles of gas and the temperature remain constant: $P_1V_1 = P_2V_2$

Substitute: $100 \text{ kPa} \times 2.00 \text{ L} = P_2 \times 3.00 \text{ L} \therefore P_2 = 66.7 \text{ kPa}$

- 2) If the number of moles of gas and the pressure remain constant: $\frac{V_1}{T_1} = \frac{V_2}{T_2}$

Substitute: $\frac{1.00 \text{ L}}{200 \text{ K}} = \frac{V_2}{400 \text{ K}} \therefore V_2 = 2.00 \text{ L}$

- 3) If the number of moles of gas and the volume remain constant: $\frac{P_1}{T_1} = \frac{P_2}{T_2}$

Substitute: $\frac{100 \text{ kPa}}{298 \text{ K}} = \frac{P_2}{398 \text{ K}} \therefore P_2 = 133.6 \text{ kPa}$

- 4) Xe has molar mass 131.3 gmol^{-1} , Ar has molar mass 39.95 gmol^{-1} , Ne has molar mass 20.18 gmol^{-1}

If equal masses of each gas are used, the sample of the gas with the smallest molar mass will contain the greatest number of moles.

Therefore, the sample of Ne will contain the greatest number of moles.

Pressure is directly proportional to the number of moles of gas; so the pressure in the Ne flask is the greatest. (Answer a)

- 5) O_2 has molar mass 32.00 gmol^{-1} , N_2 has molar mass 28.02 gmol^{-1}

If equal masses of each gas are used, the sample of the gas with the smaller molar mass will contain the greater number of moles.

Therefore, the sample of N_2 will contain the greater number of moles, and therefore molecules.

Pressure is directly proportional to the number of moles (or molecules) of gas; so the pressure in the N_2 flask is the greater than the pressure in the O_2 flask. (Answer c)

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