

# The Mole Study Guide

**Miss Combs** would your answers to 2 points after the second. Make sure to show your steps!

1) There are 2.00 moles  $\text{CO}_2$ .

a) 100g

b) 100g

c) 100g

2) There are 2.00 moles  $\text{CO}_2$ .

a) 100g

b) 100g

c) 100g

3) There are 2.00 moles  $\text{CO}_2$ .

a) 100g

b) 100g

4) There are 2.00 moles  $\text{CO}_2$ .

a) 100g

b) 100g

5) There are 2.00 moles  $\text{CO}_2$ .

a) 100g

b) 100g

Name \_\_\_\_\_ Date \_\_\_\_\_

## The Mole Study Guide

Fill in the blank boxes for the image below.



Write the empirical formula for each compound. Use subscripts to show the number of atoms.

Molecular	Empirical
$\text{C}_6\text{H}_6$	
$\text{C}_2\text{H}_4$	
$\text{C}_4\text{H}_8$	
$\text{C}_8\text{H}_{16}$	

Calculate the Molar Mass for the following Compounds.

1)  $\text{H}_2\text{O}$

2)  $\text{H}_2\text{O}_2$

3)  $\text{H}_2\text{SO}_4$

4)  $\text{C}_2\text{H}_6$

Calculate the Percent Composition for each element in the following Compound.

1)  $\text{H}_2\text{O}$

2)  $\text{H}_2\text{O}_2$

3)  $\text{H}_2\text{SO}_4$

4)  $\text{C}_2\text{H}_6$

5)  $\text{C}_2\text{H}_4$

6)  $\text{C}_2\text{H}_6$

7)  $\text{H}_2\text{O}$

8)  $\text{H}_2\text{O}_2$

Determine the Empirical and Molecular Formula for the compound given the information below.

1) A compound has the following percent composition: 40.00% C, 66.67% H, 33.33% O. The molar mass for this compound is 100.0 g/mol.

a. If you have 100 g of this compound, how many grams of H, C, and O do you have?

b. Determine the Empirical Formula for this compound.

c. Determine the Molecular Formula for this compound.

# Chapter 1the Mole Study Guide

**Morton M. Sternheim**



## Chapter 1 the Mole Study Guide:

**Study Guide for Chemical Principles** Thomas Elliott Taylor, 1979 *Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

**Study Guide to Accompany Basics for Chemistry** Martha Mackin, 1982-01-01 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

**Complete Study Guide for the Professional Engineering Examination Part 1-Fundamentals of Engineering (EIT): Fundamentals of engineering (EIT)**, 1987 *Study Guide to Chemistry and Life* John William Hill, Dorothy M. Feigl, 1978

Students Study Guide Fredeen, Pearson, 2003-07 *Student Study Guide and Solutions Manual for Gener Al Physics* Morton M. Sternheim, 1986 Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and

provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

**Chemistry, Study Guide** Bernice G. Segal, Peter S. Shenkin, 1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry chemical equilibrium and thermochemistry Worked examples now number over 300 and exercises over 1460

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*Study Guide for Introductory Chemistry* Iris Stovall, 1990

**Study Guide and Full Solutions Manual** Susan McMurry, John McMurry, 2002-08 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all end of chapter problems in the textbook

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