

Chapter 20: Oxidation-Reduction Reactions

Lesson 20.1: The Meaning of Oxidation and Reduction

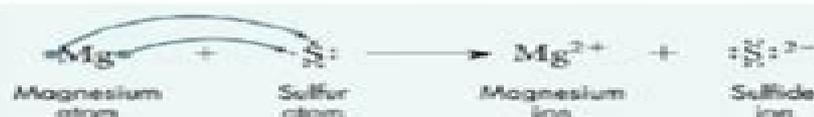
- Early chemists saw oxidation only as the combining of an element with oxygen to produce an oxide.

Figure 20.1
Oxidation of Methane
In a Bunsen burner, oxygen from the air is mixed with methane from the gas line. The methane is oxidized, and carbon dioxide and water are produced.



- *Not all* oxidation processes that use oxygen involve burning.
- A process called **reduction** is the opposite of **oxidation**.
- Oxidation and reduction always occur simultaneously.
- A substance that undergoes **oxidation gains oxygen**.
- A substance that undergoes **reduction loses oxygen**.
- No oxidation occurs without reduction, and no reduction occurs without oxidation.
- Reactions that involve the processes of oxidation and reduction are therefore called **oxidation-reduction reactions**.
- Oxidation-reduction reactions are also known as **redox reactions**.
- **Oxidation** is now defined to mean *complete or partial loss of electrons or gain of oxygen*.
- **Reduction** is now defined to mean *complete or partial gain of electrons or loss of oxygen*.

Oxidation	Reduction
Loss of electrons	Gain of electrons
Gain of oxygen	Loss of oxygen



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Table of Contents Chapter 20 Oxidation Reduction Reactions Guided Reading

1. Understanding the eBook Chapter 20 Oxidation Reduction Reactions Guided Reading
 - The Rise of Digital Reading Chapter 20 Oxidation Reduction Reactions Guided Reading
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 20 Oxidation Reduction Reactions Guided Reading
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 20 Oxidation Reduction Reactions Guided Reading
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 20 Oxidation Reduction Reactions Guided Reading

- Personalized Recommendations
 - Chapter 20 Oxidation Reduction Reactions Guided Reading User Reviews and Ratings
 - Chapter 20 Oxidation Reduction Reactions Guided Reading and Bestseller Lists
5. Accessing Chapter 20 Oxidation Reduction Reactions Guided Reading Free and Paid eBooks
 - Chapter 20 Oxidation Reduction Reactions Guided Reading Public Domain eBooks
 - Chapter 20 Oxidation Reduction Reactions Guided Reading eBook Subscription Services
 - Chapter 20 Oxidation Reduction Reactions Guided Reading Budget-Friendly Options
 6. Navigating Chapter 20 Oxidation Reduction Reactions Guided Reading eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 20 Oxidation Reduction Reactions Guided Reading Compatibility with Devices
 - Chapter 20 Oxidation Reduction Reactions Guided Reading Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 20 Oxidation Reduction Reactions Guided Reading
 - Highlighting and Note-Taking Chapter 20 Oxidation Reduction Reactions Guided Reading
 - Interactive Elements Chapter 20 Oxidation Reduction Reactions Guided Reading
 8. Staying Engaged with Chapter 20 Oxidation Reduction Reactions Guided Reading
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 20 Oxidation Reduction Reactions Guided Reading
 9. Balancing eBooks and Physical Books Chapter 20 Oxidation Reduction Reactions Guided Reading
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 20 Oxidation Reduction Reactions Guided Reading
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chapter 20 Oxidation Reduction Reactions Guided Reading
 - Setting Reading Goals Chapter 20 Oxidation Reduction Reactions Guided Reading
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chapter 20 Oxidation Reduction Reactions Guided Reading

- Fact-Checking eBook Content of Chapter 20 Oxidation Reduction Reactions Guided Reading
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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