

CHAPTER 24 REVIEW

Nuclear Chemistry

SECTION 1

Write a complete answer to the following questions in the space provided.

1. Describe the relationship between the three primary particles in the atom as well as their properties.
 - a. Masses.
 - b. Charges.
 - c. Locations.
 - d. They all have the same mass.
2. The electron has a negative charge of:
 - a. -1.6×10^{-19} coulombs.
 - b. -1.6×10^{-18} coulombs.
 - c. -1.6×10^{-17} coulombs.
 - d. -1.6×10^{-16} coulombs.
3. The number of electrons in an atom is equal to the:
 - a. number of protons.
 - b. number of neutrons.
 - c. number of electrons.
 - d. ratio of neutrons to protons.
4. What does a beta minus decay (or conversion) produce?
 - a. neutron is converted into a proton in nucleus.
 - b. proton is converted into a neutron.
 - c. neutron is gained and energy is released.
 - d. neutron is gained and energy is released.
5. The nature of a gamma-ray emission from a radioactive substance is:
 - a. smaller binding energy per nucleon.
 - b. greater binding energy per nucleon.
 - c. decreased binding energy per nucleon as the nuclei are left in a lower state.
 - d. smaller binding energy per nucleon.
6. Based on Figure 24.1 of the text, which range of E_{kin} values is characteristic of:
 - a. fast neutron fission energy per nucleon?
 - b. average kinetic energy per nucleon?
7. The number of neutrons in an atom of hydrogen-2 is _____.
8. The mass of the proton is about _____ times that of the electron.

Chapter 21 Review Nuclear Chemistry Section 1 Answer Key

M Walker

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