

**Chapter 23 Worksheet 4 (ws23.4)**  
**Nuclear Energy: Nuclear Binding Energy, Fission, and Fusion**

Einstein discovered that matter could be converted to energy (and vice-versa). The equation that expresses this mass-energy equivalency is:

$$E = mc^2 \quad (c = 3.00 \times 10^8 \text{ m/s})$$

or

$$\Delta E = (\Delta m)c^2$$

Every process that releases energy is accompanied by an equivalent loss of mass. Every process that absorbs energy results in a gain of mass. The mass changes accompanying chemical reactions are too small to measure but mass changes due to nuclear reactions can be measured using a **mass spectrometer**. The following process releases energy (how do you know?):



Thus, the mass of a nucleus is less than the sum of the masses of the protons and neutrons from which it is composed! The difference in mass is called the *mass defect* ( $\Delta m$ ):

$$\Delta m = \text{mass of protons} + \text{mass of neutrons} - \text{mass of nucleus}$$

**(Note: In some problems instead of the nuclear mass you will be given the atomic mass. In such cases, to calculate the nuclear mass you must subtract the mass of the electrons.)**

You can calculate the energy released in this process using Einstein's equation. The energy equivalent of the mass defect is called the *nuclear binding energy*,  $\Delta E$ :

$$\Delta E = (\Delta m)c^2$$

(If energy is in joules what must the unit be for mass?)

The binding energy represents the amount of energy that must be supplied to break the nucleus into its individual protons and neutrons. (Conversely, it is the energy released when the nucleus is formed from individual protons and neutrons.) For a *single nucleus*, the magnitude of the binding energy is typically between  $10^{-10}$  to  $10^{-12}$  J. (For convenience, nuclear binding energies are sometimes expressed in mega electron volts (MeV), where  $1 \text{ MeV} = 10^6 \text{ eV} = 1.60 \times 10^{-13} \text{ J}$ )

A more useful quantity for comparing the stability of nuclides is the *binding energy per nucleon*:

$$\text{Binding Energy per nucleon} = \Delta E/A$$

where  $A$  = the mass number of the isotope (number of nucleons). The greater the binding energy per nucleon, the more stable the nucleus.

# Chapter 22 Review Nuclear Chemistry Section 22 Answers

**IM Harris**



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