

Teil 2 (70% der Gesamtpunkte)



1. Die folgenden Stoffe sind in wässriger Lösung gegeben. Berechnen Sie die molare Leitfähigkeit Λ_m bei 25 °C.

Stoff	c / mol/L	κ / S/cm	Λ_m / S cm ² mol ⁻¹	Λ_m^0 / S cm ² mol ⁻¹
KCl	0,01	0,141	141	149,9
NaCl	0,01	0,106	106	126,4
HCl	0,01	0,426	426	426,2
HNO ₃	0,01	0,418	418	410,2
H ₂ SO ₄	0,01	0,859	859	800,4

Angabe: $\kappa = 0,01 \text{ mol/L} \cdot \Lambda_m$ (für univalenten Elektrolyten) bzw. $\kappa = 0,01 \text{ mol/L} \cdot z_+ z_- \Lambda_m$ (für polyvalente Elektrolyten)

2. Die molare Leitfähigkeit Λ_m von HCl bei 25 °C ist 426,2 S cm² mol⁻¹. Die molare Leitfähigkeit Λ_m^0 von HCl bei 25 °C ist 426,2 S cm² mol⁻¹.
 a) Berechnen Sie die molare Leitfähigkeit Λ_m von HCl bei 25 °C für eine 0,01 mol/L HCl-Lösung.
 b) Berechnen Sie die molare Leitfähigkeit Λ_m von NaCl bei 25 °C für eine 0,01 mol/L NaCl-Lösung.
 c) Berechnen Sie die molare Leitfähigkeit Λ_m von H₂SO₄ bei 25 °C für eine 0,01 mol/L H₂SO₄-Lösung.
 d) Berechnen Sie die molare Leitfähigkeit Λ_m von HNO₃ bei 25 °C für eine 0,01 mol/L HNO₃-Lösung.
 e) Berechnen Sie die molare Leitfähigkeit Λ_m von KCl bei 25 °C für eine 0,01 mol/L KCl-Lösung.

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 d) Berechnen Sie die molare Leitfähigkeit Λ_m von HNO₃ bei 25 °C für eine 0,01 mol/L HNO₃-Lösung.
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 e) Berechnen Sie die molare Leitfähigkeit Λ_m von KCl bei 25 °C für eine 0,01 mol/L KCl-Lösung.

Angabe: $\kappa = 0,01 \text{ mol/L} \cdot \Lambda_m$ (für univalenten Elektrolyten) bzw. $\kappa = 0,01 \text{ mol/L} \cdot z_+ z_- \Lambda_m$ (für polyvalente Elektrolyten)

Chapter 1 physics Test

Mike Jess



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