

CHAPTER 14 Acids and bases in the environment

Chapter review

Name: _____

MULTIPLE CHOICE QUESTIONS

- Which of the following is the best definition of an acid?
 - An acid is an electron acceptor.
 - An acid is a substance that tastes sour.
 - An acid is a substance that donates H^+ .
 - An acid is a substance that accepts H^+ .
- Consider the following equation and state the species which are acting as bases in this reaction.



- $HCN(aq)$ and $H_2O(l)$
 - $HCN(aq)$ and $H_3O^+(aq)$
 - $HCN(aq)$ and $CN^-(aq)$
 - $H_2O(l)$ and $CN^-(aq)$
- D
- The sourness of food is a reasonable guide to its acidity, since sourness increases with increasing acidity. If food A is more sour than food B, then it is likely that:
 - A has a higher pH than solution B
 - A has more $OH^-(aq)$ ions than B
 - B is a stronger acid than A
 - A has a higher concentration of $H_3O^+(aq)$ ions than B.
 - Hydrogen sulfide, H_2S , is a gas at room temperature and dissolves in water. When this solution is tested, it conducts electricity and turns litmus solution red. The solution is acidic because:
 - the H_2S has gained a proton
 - the H_2O has gained a proton
 - the H_2S and H_2O have gained protons
 - the H_2S and H_2O have lost protons.
 - The conjugate base of the species $H_2PO_4^-(aq)$ is:
 - $H_3PO_4(aq)$
 - $HPO_4^{2-}(aq)$
 - $H_2PO_4(aq)$

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General Chemistry for Engineers Jeffrey Gaffney, Nancy Marley, 2017-11-13 *General Chemistry for Engineers* explores the key areas of chemistry needed for engineers This book develops material from the basics to more advanced areas in a systematic fashion As the material is presented case studies relevant to engineering are included that demonstrate the strong link between chemistry and the various areas of engineering Serves as a unique chemistry reference source for professional engineers Provides the chemistry principles required by various engineering disciplines Begins with an atoms first approach building from the simple to the more complex chemical concepts Includes engineering case studies connecting chemical principles to solving actual engineering problems Links chemistry to contemporary issues related to the interface between chemistry and engineering practices *Study Questions and Problems in Inorganic Chemistry* Alexander Silverman, 1924

A Problem-Solving Approach to Aquatic Chemistry James N. Jensen, 2022-12-20 *A Problem Solving Approach to Aquatic Chemistry* Enables civil and environmental engineers to understand the theory and application of aquatic equilibrium chemistry The second edition of *A Problem Solving Approach to Aquatic Chemistry* provides a detailed introduction to aquatic equilibrium chemistry calculation methods for systems at equilibrium applications of aquatic chemistry and chemical kinetics The text directly addresses two required ABET program outcomes in environmental engineering chemistry including stoichiometry equilibrium and kinetics and material and energy balances fate and transport of substances in and between air water and soil phases The book is very student centered with each chapter beginning with an introduction and ending with a summary that reviews the chapter s main points To aid in reader comprehension important terms are defined in context and key ideas are summarized Many thought provoking discussion questions worked examples and end of chapter problems are also included Each part of the text begins with a case study a portion of which is addressed in each subsequent chapter illustrating the principles of that chapter In addition each chapter has an Historical Note exploring connections with the people and cultures connected to topics in the text *A Problem Solving Approach to Aquatic Chemistry* includes Fundamental concepts such as concentration units thermodynamic basis of equilibrium and manipulating equilibria Solutions of chemical equilibrium problems including setting up the problems and algebraic graphical and computer solution techniques Acid base equilibria including the concepts of acids and bases titrations and alkalinity and acidity Complexation including metals ligands equilibrium calculations with complexes and applications of complexation chemistry Oxidation reduction equilibria including equilibrium calculations graphical approaches and applications Gas liquid and solid liquid equilibrium with expanded coverage of the effects of global climate change Other topics including chemical kinetics of aquatic systems surface chemistry and integrative case studies For advanced senior undergraduates and first year graduate students in environmental engineering courses *A Problem Solving Approach to Aquatic Chemistry* serves as an invaluable learning resource on the topic with a variety of helpful learning elements included throughout to ensure information retention and the

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Instructor's Resource Manual Susan M. Young, 2006 Chemistry Oxtoby, 1994 **Practical Chemistry** Lyman

Churchill Newell,1922 **The Westminster review** ,1884 *Regulations for the Preliminary Education of Elementary School Teachers* Great Britain. Board of Education,1909 **Introduction to chemistry** Martha J. Gilleland,Robert C. Pfaff,1986 **First Year Chemistry Students' Conceptions of Acid/base Chemistry** Sally Diane Rupert,2001 Annual Report - Dept. of Education Saskatchewan. Department of Education,1913 **Annual Report** ,1911 *Annual Report* Saskatchewan. Department of Education,1913 *Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson* Norman E. Griswold,Henry Fuller Holtzclaw,1988

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