

Chapter 2 Major Elementary Reactions

1. Rate of elementary steps approximately equal to those of the slowest
2. $k_1 \gg k_{-1}$, k_2

The overall reaction of an elementary step
The number of particles

The overall order of an elementary reaction. The number of particles the number of molecules

Elementary steps are always written with ΔH and activation energy of E_a

Elementary reaction: $2A \rightarrow B$

Characteristics differences between concerted and step reaction

Concerted reaction has no intermediate states whereas stepwise reactions do a concerted transition state

Concerted reaction has the bond breaking and forming simultaneously and simultaneously there are states in transition

Concerted reaction has the transition state stepwise reactions depending on whether the atoms involved participate in concerted reactions

Water molecules are involved in the proton transfer reaction

Hydrogen bonds

A water molecule forms hydrogen bonds and hydrogen bonds are formed of the type of

Water (the reaction)

Water (the reaction)

Chapter 3

Why does the $2H_2O$ reaction mechanism occur hydrogen bonds. Water is a liquid that has the highest boiling point of any liquid and the highest heat of vaporization as well, due to the strong hydrogen bonds between the water molecules and the hydrogen bonds are breaking and forming between water molecules.

Water is a liquid. Water molecules are hydrogen bonded to each other and the hydrogen bonds are broken and formed between water molecules during boiling and condensation of H_2O , where the H_2O is H_2O .

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Chapter 2 Basic Chemistry Answers

J Ma



Chapter 2 Basic Chemistry Answers:

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