

CHAPTER 2- ACIDS AND BASES

Questions And Answers 2023-2024

According to the Bronsted-Lowry definition, a base is a proton _____, while an acid is a proton _____. -
Correct Answer-acceptor; donor

According to the Bronsted-Lowry definition, an acid must contain which of the following? - Correct Answer-
A hydrogen atom

A Bronsted-Lowry base must have an "available" electron pair to form a bond to H^+ . Which of the following structural features allow an electron pair to be used in this way? - Correct Answer-- One or more electron pairs in π bonds.

- One or more lone pairs in the structure.

Which of the following Bronsted-Lowry acid/base classifications are correct? - Correct Answer-- $(CH_3)_3N$ is a base.

- CH_3COOH is an acid.

Select all statements that accurately describe Bronsted-Lowry acid-base reactions. - Correct Answer--
Acid-base reactions are also known as proton transfer reactions.

- One bond is formed in an acid-base reaction.

Match each term with the correct Bronsted-Lowry definition.

- Acid:

- Base: - Correct Answer-- proton (H^+) donor

- proton (H^+) acceptor

When identifying Bronsted-Lowry acids and bases in an organic reaction, if only one species has a lone pair or π bond it must be the _____, while if only one species contains an available hydrogen atom it must be the _____. - Correct Answer-base, acid

*Reason: A species with a lone pair can readily accept a proton and is therefore defined as a base.

Which of the following statements correctly describe Bronsted-Lowry acids? - Correct Answer-- A Bronsted-Lowry acid must contain a H atom.

- A Bronsted-Lowry acid is a proton donor.

Chapter 2 acid And Bases Answers

William H. Brown, Thomas Poon



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