

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____	lattice energy	_____	a. ionic bond
_____	bond energy	_____	b. covalent bond
_____	enthalpy of vaporization	_____	c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____	ionic	a. H—F	_____	nonpolar covalent	d. H—H
_____	ionic	b. Na—Cl	_____	polar covalent	e. H—C
_____	polar covalent	c. H—O	_____	polar covalent	f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond. Increased polarity in H₂O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

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Guillaume Favre



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In some sort of driven by information and connectivity, the energy of words has be more evident than ever. They have the ability to inspire, provoke, and ignite change. Such is the essence of the book **Chapter 8 Covalent Bonding Pearson Worksheet Answer Key**, a literary masterpiece that delves deep to the significance of words and their effect on our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book is key themes, examine its writing style, and analyze its overall affect readers.

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