

Acids and Bases Worksheet

NAME _____

1. Name the following compounds as acids and circle which are weak acids:
 a) HNO_3 b) H_2SO_4 c) HF d) H_2CO_3 e) $\text{HC}_2\text{H}_3\text{O}_2$ f) H_3PO_4
2. Write the formulas for the following bases:
 a) calcium hydroxide b) aluminum hydroxide c) ammonium hydroxide d) lithium hydroxide
3. Complete the following table:

	Acids	Bases
Definition		
Electrolyte in aqueous solution?		
pH		
Taste		
Litmus turns		
Which reacts with metals to form H_2 gas?		

Conjugate acid-base pairs

4. Identify the acid, the base, the conjugate acid and the conjugate base in the reaction:
 $\text{NH}_3(\text{g}) + \text{HOH}(\text{l}) \rightleftharpoons \text{NH}_4^+(\text{aq}) + \text{OH}^-(\text{aq})$
5. Identify the acid, the base, the conjugate acid and the conjugate base in the reaction:
 $\text{HCO}_3^-(\text{aq}) + \text{HOH}(\text{l}) \rightleftharpoons \text{CO}_3^{2-}(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$
6. An acid will form the conjugate _____ and a base will form the conjugate _____.

Neutralization and Titration.

7. Classify the following acids as monoprotic, diprotic or triprotic:
 a) H_2SO_4 b) H_3PO_4 c) HCl d) $\text{HC}_2\text{H}_3\text{O}_2$ or CH_3COOH (same thing)
8. What is the concentration of an unknown NaOH solution if took an average of 15.4mL of 0.100N HCl solution to neutralize 10.0mL of the NaOH solution?
9. From the given data, what is the molarity and the normality of the HF solution if the KOH solution was 0.50M?

	Acid, HF	Base, KOH
Initial Volume, mL	5.5	14.9
Final Volume, mL	25.9	24.4

$\text{pH} = -\log[\text{H}_3\text{O}^+]$ To find pH if given $[\text{H}_3\text{O}^+]$, enter the $[\text{H}_3\text{O}^+]$, press **Log** & then the **+/−** button
 To find $[\text{H}_3\text{O}^+]$ if given pH, enter negative pH, then **2nd** function **Log**

10. A solution has a $[\text{H}_3\text{O}^+]$ of 1×10^{-7} M, what is its pH? _____ Is this an acid, base or is it neutral?

TURN THE PAPER OVER

11. What is the pH of a solution if the $[\text{H}^+] = 7.2 \times 10^{-9}\text{M}$? _____ Is this an acid or base?

Chapter Acids And Bases Answer Key

B Lingard



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Table of Contents Chapter Acids And Bases Answer Key

1. Understanding the eBook Chapter Acids And Bases Answer Key
 - The Rise of Digital Reading Chapter Acids And Bases Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Acids And Bases Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter Acids And Bases Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter Acids And Bases Answer Key
 - Personalized Recommendations
 - Chapter Acids And Bases Answer Key User Reviews and Ratings
 - Chapter Acids And Bases Answer Key and Bestseller Lists
5. Accessing Chapter Acids And Bases Answer Key Free and Paid eBooks
 - Chapter Acids And Bases Answer Key Public Domain eBooks
 - Chapter Acids And Bases Answer Key eBook Subscription Services
 - Chapter Acids And Bases Answer Key Budget-Friendly Options

6. Navigating Chapter Acids And Bases Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter Acids And Bases Answer Key Compatibility with Devices
 - Chapter Acids And Bases Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter Acids And Bases Answer Key
 - Highlighting and Note-Taking Chapter Acids And Bases Answer Key
 - Interactive Elements Chapter Acids And Bases Answer Key
8. Staying Engaged with Chapter Acids And Bases Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter Acids And Bases Answer Key
9. Balancing eBooks and Physical Books Chapter Acids And Bases Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter Acids And Bases Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter Acids And Bases Answer Key
 - Setting Reading Goals Chapter Acids And Bases Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter Acids And Bases Answer Key
 - Fact-Checking eBook Content of Chapter Acids And Bases Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

-
- Interactive and Gamified eBooks

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