

## Chapter 4 Practice Problems

1. How many grams of  $\text{Li}_3\text{N}$  can be formed from 1.75 moles of  $\text{Li}$ ? Assume an excess of nitrogen.



$$1.75 \text{ mol Li} \left( \frac{2 \text{ mol Li}_3\text{N}}{6 \text{ mol Li}} \right) \left( \frac{35.01 \text{ g Li}_3\text{N}}{1 \text{ mol}} \right) = \boxed{172.535 \text{ g Li}_3\text{N}}$$

2. How many moles of oxygen are formed when 58.6 g of  $\text{KNO}_3$  decomposes according to the following reaction? The molar mass of  $\text{KNO}_3$  is 101.11 g/mol.



$$58.6 \text{ g} \left( \frac{1 \text{ mol}}{101.11 \text{ g}} \right) \left( \frac{5 \text{ mol O}_2}{4 \text{ mol KNO}_3} \right) = \boxed{0.724 \text{ mol O}_2}$$

3. How many moles of nitrogen are formed when 58.6 g of  $\text{KNO}_3$  decomposes according to the following reaction? The molar mass of  $\text{KNO}_3$  is 101.11 g/mol.



$$58.6 \text{ g} \left( \frac{1 \text{ mol}}{101.11 \text{ g}} \right) \left( \frac{2 \text{ mol N}_2}{4 \text{ mol KNO}_3} \right) = \boxed{0.290 \text{ mol N}_2}$$

4. A 14.01 g sample of  $\text{N}_2$  reacts with 3.02 g of  $\text{H}_2$  to form ammonia ( $\text{NH}_3$ ). If ammonia is the only product, what mass of ammonia is formed?

$$14.01 + 1.008(3) = 17.034$$



$$14.01 \text{ g} \left( \frac{1 \text{ mol}}{28.02} \right) \left( \frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} \right) = 1.00 \text{ mol}$$

$$3.02 \text{ g} \left( \frac{1 \text{ mol}}{2.016} \right) \left( \frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} \right) = 0.999 \text{ mol} \left( \frac{17.034}{1 \text{ mol}} \right) = \boxed{17.02 \text{ g NH}_3}$$

5. Consider the following balanced reaction. How many grams of water are required to form 75.9 g of  $\text{HNO}_3$ ? Assume that there is excess  $\text{NO}_2$  present. The molar masses are as follows:  $\text{H}_2\text{O} = 18.02$  g/mol,  $\text{HNO}_3 = 63.02$  g/mol.



$$75.9 \text{ g} \left( \frac{1 \text{ mol}}{63.02 \text{ g}} \right) \left( \frac{1 \text{ mol H}_2\text{O}}{2 \text{ mol HNO}_3} \right) \left( \frac{18.02 \text{ g}}{1 \text{ mol}} \right) = \boxed{10.9 \text{ g H}_2\text{O}}$$

# Chapter Practice Problems

**Richard Bailey**



**Chapter Practice Problems:**

## **Chapter Practice Problems** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Chapter Practice Problems**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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