

# Chm 113 Final Study Guide Questions & Answers 100% Correct.

What is the  $[H^+]$  of a solution with a pH of 3.78? - ANSWER  
 $pH = 3.78$   
 $[H^+] = 10^{-3.78}$   
 $= 1.66 \times 10^{-4}$

If HX represents the formula of a weak acid, what is the formula of the conjugate base of HX? - ANSWER  
If HX weak acid then the formula of the conjugate base of X-

Let HX represent the formula of a weak acid. Write a balanced chemical equation for the reaction of this weak acid with the strong base NaOH? You may omit any spectator ions if you wish. - ANSWER  
 $HX + NaOH \rightarrow NaX + H_2O$   
For the equation you wrote above, what two terms are considered/assumed equal? - ANSWER  
 $[HA] = [A^-]$

Regarding today's experiment, what is the weak acid we are going to titrate and which strong base is used to titrate this weak acid? - ANSWER  
Weak Acid: Acetic Acid  
Strong Base: NaOH

If a weak acid, HX, is titrated with a strong base such as NaOH, will the equivalence point in the titration occur below pH 7.0, at pH 7.0 or above pH 7.0? - ANSWER  
the equivalence point in the titration will be somewhat above pH 7.0

A student is presented with a 0.65 M solution of an unknown weak acid. The student measures the pH of that solution and determines a value of 3.78 for the unknown solution. What, then, is the  $K_a$  of this weak acid? - ANSWER  
 $K_a = \frac{(1.66 \times 10^{-4})^2}{.65 - 1.66 \times 10^{-4}} = 4.2 \times 10^{-8}$

What is the acid dissociation constant ( $K_a$ )? - ANSWER  
 $K_a = \frac{[H_3O^+][A^-]}{[HA]}$

What is the pH of a weak acid whose  $K_a$  is  $6.03 \times 10^{-6}$ , and the ratio of  $[A^-]/[HA]$  is 3. - ANSWER  
 $6.03 \times 10^{-6} = \frac{[H_3O^+][9]/3}{[HA]}$   
 $pH = 5.69$

# 113 Study Guide Answers 129346

**Rachel S Tattersall**



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## **Table of Contents 113 Study Guide Answers 129346**

1. Understanding the eBook 113 Study Guide Answers 129346
  - The Rise of Digital Reading 113 Study Guide Answers 129346
  - Advantages of eBooks Over Traditional Books
2. Identifying 113 Study Guide Answers 129346
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 113 Study Guide Answers 129346
  - User-Friendly Interface
4. Exploring eBook Recommendations from 113 Study Guide Answers 129346
  - Personalized Recommendations
  - 113 Study Guide Answers 129346 User Reviews and Ratings
  - 113 Study Guide Answers 129346 and Bestseller Lists
5. Accessing 113 Study Guide Answers 129346 Free and Paid eBooks
  - 113 Study Guide Answers 129346 Public Domain eBooks
  - 113 Study Guide Answers 129346 eBook Subscription Services
  - 113 Study Guide Answers 129346 Budget-Friendly Options
6. Navigating 113 Study Guide Answers 129346 eBook Formats

- ePub, PDF, MOBI, and More
  - 113 Study Guide Answers 129346 Compatibility with Devices
  - 113 Study Guide Answers 129346 Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of 113 Study Guide Answers 129346
  - Highlighting and Note-Taking 113 Study Guide Answers 129346
  - Interactive Elements 113 Study Guide Answers 129346
8. Staying Engaged with 113 Study Guide Answers 129346
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 113 Study Guide Answers 129346
9. Balancing eBooks and Physical Books 113 Study Guide Answers 129346
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection 113 Study Guide Answers 129346
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 113 Study Guide Answers 129346
- Setting Reading Goals 113 Study Guide Answers 129346
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 113 Study Guide Answers 129346
- Fact-Checking eBook Content of 113 Study Guide Answers 129346
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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